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In Situ Plasmonic Nanospectroscopy of the CO Oxidation Reaction over Single Pt Nanoparticles

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Supporting Information

ABSTRACT: The ongoing quest to develop single-particle methods for the in situ study of heterogeneous catalysts is driven by the fact that heterogeneity in terms of size, shape, grain structure, and composition is a general feature among nanoparticles in an ensemble. This heterogeneity hampers the generation of a deeper understanding for how these parameters affect catalytic properties. Here we present a solution that in a single benchtop experimental setup combines single-particle plasmonic nanospectroscopy with mass spectrometry for gas phase catalysis under reaction conditions at high temperature. We measure changes in the surface state of polycrystalline platinum model catalyst particles in the 70 nm size range and the corresponding bistable kinetics during the carbon monoxide oxidation reaction via the peak shift of the dark-field scattering spectrum of a closely adjacent plasmonic nanoantenna sensor and compare these changes with the total reaction rate measured by the mass spectrometer from an ensemble of nominally identical particles. We find that the reaction kinetics of simultaneously measured individual Pt model catalysts are dictated by the grain structure and that the superposition of the individual nanoparticle response can account for the significant broadening observed in the corresponding nanoparticle ensemble data. In a wider perspective our work enables in situ plasmonic nanospectroscopy in controlled gas environments at high temperature to investigate the role of the surface state on transition metal catalysts during reaction and of processes such as alloying or surface segregation in situ at the single-nanoparticle level for model catalysts in the few tens to hundreds of nanometer size range.

KEYWORDS: single-particle catalysis, single particle, plasmonic nanospectroscopy, CO oxidation, bistable kinetics, dark-field scattering spectroscopy, quadrupole mass spectrometry

Studying individual nanoparticles is of high relevance in heterogeneous catalysis,1−4 where they are widely used and where polydispersity in terms of size, shape, and grain structure is a general feature among them. This heterogeneity hampers the generation of a deeper understanding for how these structural parameters affect catalytic activity since they, together with electronic and spillover interactions with the support,5−8 directly control the catalytic performance. Assessing the state, activity, and selectivity of individual nanoparticles thus has significant potential to contribute to the development of efficient catalyst materials. Therefore, the characterization of single nanoparticles at in situ conditions is a major tour de force in catalysis, and significant efforts are invested in the development of the required experimental techniques.

To this end, plasmonic nanospectroscopy is an experimental concept that employs metal nanoparticles capable of manipulating light at the nanoscale, via electron oscillations known as localized surface plasmon resonance (LSPR).9−11 Such plasmonic nanoantennas have been successfully used as nanoscopic probes of various processes including phase transitions,11 biomolecule interactions and sensing,12−14 metal hydride formation,15 gas- and chemosensing10,16 and catalytic reactions.17−21 In catalysis applications they can report directly on the catalyst nanoparticle surface or bulk state due to the intrinsically very high sensitivity of LSPR to surface and bulk changes.11 One of the most appealing assets of the
The plasmonic nanospectroscopy concept, when projected onto catalysis research, is the possibility to address individual nanoparticles in the 10–100s nm size regime, which is accessible to the enhanced precision of top-down nanofabrication that enables the preparation of controlled and precisely tunable model systems.\textsuperscript{4,22–25} Notably, this single-nanoparticle resolution comes without principle restrictions on the surrounding medium; that is, both liquid and gas phase environments are accessible at ambient pressure or above.\textsuperscript{6,13,19,26–31} Moreover, multiple individual nanoparticles can be addressed simultaneously using concepts such as hyperspectral imaging and derivatives.\textsuperscript{32–35} This potentially offers similar insights as recently obtained by in situ X-ray absorption spectroscopy at a beamline\textsuperscript{6,7} but with benchtop optical microscopy instrumentation for optical dark-field scattering spectroscopy.\textsuperscript{36}

To date, the plasmonic nanospectroscopy concept has been successfully applied to study catalytic processes at the single-particle level at \textit{in situ} conditions in the liquid phase, as introduced in the seminal paper by Novo \textit{et al.}\textsuperscript{19} Subsequent studies by various groups further diversified the concept\textsuperscript{37} to, for instance, investigate the photocatalytic decomposition of lactic acid,\textsuperscript{38} Au nanoparticle-catalyzed redox reactions between glucose and \textit{O}_2,\textsuperscript{39} hydrogenation of chemisorbed 4-NTP,\textsuperscript{39} electron transfer rates from different nanoparticle facets,\textsuperscript{40} spectroelectrochemistry,\textsuperscript{41} and spillover effects to an oxide support.\textsuperscript{6} However, while elegantly demonstrating the single-particle capabilities of plasmonic nanospectroscopy in catalysis, all these studies have been carried out in the liquid phase and at or close to room temperature. Thus, they leave unaddressed the important field of gas phase heterogeneous catalysis, which typically takes place at several hundred Kelvin at atmospheric pressure or above. Furthermore, all these investigations have in common that they solely rely on the plasmonic signal, that is, the spectral shift of the resonance peak as the readout. In other words, only the state of the catalyst nanoparticle itself (\textit{e.g.}, charge or oxidation state)\textsuperscript{19,40} or of the surrounding medium (\textit{e.g.}, \textit{via} spillover)\textsuperscript{6} is reported, while the reaction products remain unanalyzed.

As a first step toward overcoming these limitations, we report an experimental concept and the corresponding setup that combines \textit{in situ} single-particle plasmonic nanospectroscopy of catalyst nanoparticles with a gas phase catalytic flow reactor operating at up to 623 K and equipped with a quadrupole mass spectrometer (QMS) for the simultaneous quantitative analysis of reaction products from an adjacent ensemble (on the order of 10^9 particles) of nominally identical nanoparticles prepared by nanofabrication (Figure 1). This design enables the direct investigation of ensemble averaging effects in a single experiment since individual nanoparticles (LSPR) and a corresponding ensemble (QMS) can be probed simultaneously in the same experiment. To this end, we also note that for the detection of bioanalytes in the liquid phase both SPR\textsuperscript{32} and LSPR\textsuperscript{35} sensing setups have been combined with mass spectrometry, however, not simultaneously as done here.

As the second step we introduce an Au@SiO_2–Pt nanostructure design according to the indirect plasmonicsensing\textsuperscript{17,44} scheme, where a nanofabricated single Au plasmonic antenna observer, which is completely encapsulated in SiO_2 and closely adjacent to the active single Pt catalyst nanoparticle, boosts the total light-scattering signal of the structure and serves as optical transducer in the plasmonic nanospectroscopy single-nanoparticle readout. This arrangement is necessary since most of the transition metal catalysts are poor light scatterers,\textsuperscript{36} and it has the potential to enable studies of individual catalyst particles in the sub-10 nm size regime most relevant to industrial catalysis.\textsuperscript{28,46,47} Here, we use nanofabricated \textit{20×70 nm polycrystalline} Pt nanoparticles as our model system to deliberately create sizable particle heterogeneity in terms of grain structure, defects, and surface faceting to demonstrate how such parameters impact reaction kinetics at the single-nanoparticle level.

As the third step and to demonstrate and benchmark our setup using these Au@SiO_2–Pt structures, we investigated the kinetic phase transition phenomenon and the corresponding bistable kinetics that have been reported for the carbon monoxide (CO) oxidation reaction both on single-crystal surfaces\textsuperscript{48} and on nanofabricated Pt model catalysts.\textsuperscript{22,48,49} As the main results obtained at atmospheric pressure conditions, we found that the reaction kinetics of simultaneously measured individual model catalyst nanoparticles are remarkably different and critically depend on their grain structure in terms of abundance of grains and of the corresponding surface faceting and that the superposition of the individual nanoparticle response can account for the significant broadening observed in the corresponding nanoparticle ensemble data.

### RESULTS AND DISCUSSION

Our experimental setup (Figure 1; for a more technical drawing refer to Figure S1 in the Supporting Information, SI) comprises an upright optical microscope operated in dark-field epi-illumination mode. It is connected to a spectrometer equipped with a CCD camera, which is used to collect scattered light from (individual) plasmonic nanostructures via
a long working-distance 50× objective and thus enables both ensemble and single-particle plasmonic nanoplasmonic spectroscopy readout from a tailor-made catalytic reactor described below. Plasmonic nanoplasmonics relies on the fact that the LSPR frequency of a metal nanoparticle is very sensitive to minute changes on its surface or on a second nanoparticle in its close vicinity. Using dark-field scattering spectroscopy, the induced change in LSPR frequency can be efficiently detected as a spectral shift of the peak maximum in the light-scattering spectrum of a single nanoparticle, with resolution high enough to detect single molecules. In our setup focusing on catalysis applications, the nanofabricated plasmonic/catalytic sample is mounted on an inert flat ceramic carbon heater installed inside a glass flow reactor tube with 400 mm diameter, which we have equipped with a 1 in., 200 μm thin flat optical window to facilitate dark-field scattering spectroscopy readout from inside the reactor. It is also equipped with a gas inlet connected to mass-flow controllers for accurate control of reactant concentration. An active feedback loop controls the temperature of the sample up to 623 K via a thermocouple and a temperature controller. To facilitate the QMS readout directly from the nanofabricated sample surface, we use a glass capillary “sniffer”, which is mounted inside the flow reactor. Its position can be accurately controlled via an x−y−z micrometer stage connected to the reactor by a stainless steel bellow. Via a stainless steel tube, the sniffer is further connected to a UHV chamber, on which the QMS is mounted. The opening of the glass capillary sniffer toward the sample is tuned to 5 μm using the method introduced by Kasemo. In this way it effectively acts as orifice leak for local, fast-response gas sampling using the QMS by ensuring the necessary pressure drop from 1 atm inside the reactor to below the maximal operation pressure of the QMS, which is on the order of 10⁻⁶ mbar.

To initially validate the function of the setup, we benchmarked it with our earlier work on the kinetic phase transition that occurs during the hydrogen oxidation reaction over a Pt nanoparticle ensemble model catalyst due to bistable kinetics, which we had studied when introducing the indirect nanoplasmonic sensing concept. The corresponding analysis of a very similar set of experiments using our combined plasmonic nanoplasmonic spectroscopy and QMS setup presented here is summarized and discussed in detail in the SI and Figures S2 and S3. As the first key result, it reproduces the insights obtained in our previous study using a traditional quartz-tube flow reactor setup and simple optical transmittance measurements. As the second key result, it demonstrates the anticipated complementarity of the optical and the QMS signals, that is, that the QMS reports the catalyst activity of the nanofabricated model catalyst ensemble and that the plasmonic nanoplasmonic optical signal directly reports the catalyst surface state, as identified by the observed coincidence of the highest reaction rate and change of the catalyst surface state at the kinetic phase transition (Figure S2c).

To enable single-particle plasmonic nanoplasmonic spectroscopy from a Pt model catalyst, which by itself is a weak light scatterer, we want to place it closely adjacent to an inert plasmonic nanoantenna “observer” in order to enhance the total scattering cross-section of the system. However, in contrast to the earlier implementations of this concept, here we also have to consider the high operating temperature of our system and thus develop a means to spatially separate the catalyst from the Au nanoantenna to prevent alloy or intermetallic phase formation between the two. For this purpose, we further tailored our hole-mask colloidal lithography nanofabrication approach (see Methods for details), to enable the chemical vapor deposition of a thin SiO₂ layer through the nanofabrication mask that encapsulates the entire Au nanoantenna before the growth of the Pt catalyst nanoparticle (Figure 2a). In this way, due to the self-alignment, it becomes possible to grow the catalyst exclusively on top of the nanoantenna sensor, while simultaneously still completely encapsulating the underlying Au nanoantenna. This results in an Au@SiO₂−Pt hybrid nanostructure with combined sensing (via the Au nanoantenna) and catalytic (via the Pt) function (Figure 2b−d). Furthermore, since different types of separating layers, as well as catalyst particle materials, can be grown in this way, our approach is generic and can be easily expanded to other catalyst systems to tailor the catalyst formulation in a modular fashion.

To test the thermal and chemical stability of the Au@SiO₂−Pt nanostructures used here, we thermally annealed them at 623 K in 3% H₂ + 3% O₂ for 2 h and exposed them to reaction environment for the CO-oxidation reaction (3% CO + 3% O₂) at 623 K for 0.5 h and in Ar carrier gas. The corresponding scanning electron microscopy (SEM) analysis reveals their structural integrity, as well as the anticipated recrystallization of the Pt (Figure 2e−g). Furthermore, comparing a representative single-particle scattering spectrum of such a nanoarchitecture after the thermal and reaction treatment (Figure 2h) with a corresponding difference-time-domain (FDTD) simulation (Figure 2i depicts the used simulation scheme) reveals good agreement and thus further corroborates both the integrity of the nanostructure and the anticipated coupling between the two metal elements, essential for the indirect sensing principle.

To further characterize the Pt model catalyst nanoparticles, we performed transmission electron microscopy (TEM) analysis. However, since the entire Au@SiO₂−Pt hybrid structure would be too thick for TEM imaging, as well as would lead to convoluted images due to the stacked Au and Pt particles, we nanofabricated analogous Pt model catalyst nanoparticles (i.e., without Au underneath) with the same size on a TEM membrane and treated them in the same way as the real sample at the reaction conditions described above. The corresponding TEM images of representative Pt nanoparticles reveal the formation of mainly polycrystals with various numbers of grains (Figure 3a). Further resorting to our earlier detailed characterization of the grain structure in equivalently nanofabricated Pd nanoparticles using transmission Kikuchi diffraction (TKD), we conclude that each nanoparticle has its distinct and unique grain structure with different grain orientation (cf. Figure 5 in ref 35) and multiple surface facets exposed by the different grains (cf. Figure S17 in ref 35). Hence they fulfill the desired criterion to serve as single-particle model systems for structurally distinctly different catalyst nanoparticles.

As the final characterization step of the Au@SiO₂−Pt hybrid nanostructures to ensure that the SiO₂ layer indeed encapsulates the Au nanoantenna and prevents direct contact between Au and Pt elements even at reaction conditions, we carried out X-ray photoelectron spectroscopy (XPS) analysis (Figure 3b). It reveals that the distinct characteristic Pt 4f peaks are preserved even after the reaction treatment and thus corroborates that no mixing with the Au nanoantenna has occurred.
To generally characterize the catalytic properties of the Au@SiO$_2$–Pt hybrid nanostructures, we first performed ensemble-type experiments using the oxidation of carbon monoxide (CO + 1/2 O$_2$ → CO$_2$) over Pt as the model reaction. Specifically, we put our focus on the phenomenon of kinetic phase transitions and the corresponding reported kinetic bistabilities, that is, the existence of two stable kinetic regimes that may coexist for a given set of reaction conditions, which has been reported for the CO oxidation reaction both on nanofabricated Pt model catalysts and on single-crystal surfaces. In the present case, as illustrated schematically in Figure 4, this means that at low relative CO concentration, $\alpha_{\text{CO}} = [\text{CO}] / ([\text{CO}] + 1/2 \text{O}_2)$ decreases as the CO concentration increases, leading to a kinetic phase transition to a state where CO predominantly covers the surface. This behavior is consistent with the observed changes in the Pt catalyst dimensions due to recrystallization during thermal treatment.
down-sweeps in both the QMS and the agreement with Joha
constant at 9% in Ar carrier gas. Simultaneously, we reaction rate obtained by the QMS and expressed as CO2 α as predicted by theory, 49 we also observe increased or higher abundance of defects on smaller particles and (2 bistability phenomenon using molecular beam experiments "fl O to form CO2. This asymmetry in terms of poisoning in the signal exhibits a distinct change in trend. Since the displayed in Figure 5 b are nonlinear in magnitude, this increasing CO desorption rate, which ultimately eliminates the region is narrowing and eventually disappears due to the condition. Typically, at higher temperatures, the coexistence of two stable kinetic states and thus hysteresis in obtained reaction rates, depending on the initial surface condition. Typically, at higher temperatures, the coexistence region is narrowing and eventually disappears due to the increasing CO desorption rate, which ultimately eliminates the poisoning effect. 23,49

To explore this phenomenon with our setup and for the Au@SiO2−Pt model catalyst, we varied αCO in the reactant flow from a CO-rich to an O2-rich condition and back again, while keeping the total reactant concentration ([CO] + [O2]) constant at 9% in Ar carrier gas. Simultaneously, we continuously recorded both the CO2 partial pressure in the reactor via the QMS and the plasmonic nanospectroscopy peak position signal, λ, via the dark-field scattering spectroscopy readout from a sample area comprising ca. 5000 Au@SiO2−Pt nanostructures (Figure 5a). The sample temperature was set to 503 K. A complete αCO up- and down-sweep together with the corresponding QMS and λ response are summarized in Figure 5b. We observe a maximum in the reaction rate at a critical reactant mixture αCOcr = 0.04−0.05, at which the optical λ signal exhibits a distinct change in trend. Since the αCO steps displayed in Figure 5b are nonlinear in magnitude, this becomes clearer in Figure 5c, where the QMS and plasmonic nanospectroscopy data (now expressed as the shift of the LSPR scattering peak position, Δλ, with respect to the first taken data point at t = 0) are plotted as a linear function of αCO. Both for the αCO up- and down-sweep, Δλ is essentially constant for αCO < αCOcr and then rapidly spectrally blue-shifts for αCO > αCOcr to reach a steady-state value beyond αCO ≈ 0.2. The reaction rate obtained by the QMS and expressed as CO2 partial pressure in the chamber exhibits a distinct maximum that coincides with the onset of the blue-shift of Δλ.

It is now interesting to further discuss these data in the context of the bistable reaction kinetics introduced above. To this end, first, we observe hysteresis between αCO up- and down-sweeps in both the QMS and the Δλ response, in good agreement with Johánek et al. 22 who investigated the bistability phenomenon using molecular beam experiments on nanostructured Pd model catalysts over a wide size range (2−500 nm). They found a distinct particle size dependence of the hysteresis as a consequence of the interplay between a higher abundance of defects on smaller particles and fluctuations between the two kinetic reaction regimes. Second, as predicted by theory, 49 we also observe increased or decreased hysteresis width at lower and higher reaction temperature, respectively (Figure S4), in both the QMS and plasmonic nanospectroscopy response. This is the consequence of CO poisoning being less severe at higher temperature due to enhanced CO desorption. Third, we notice the excellent agreement in global trend between QMS signal and Δλ response in the CO-rich surface regime and the contrasting essentially constant Δλ signal in the O-rich surface regime, despite a significant change in reaction rate. This indicates that the plasmonic nanospectroscopy signal, Δλ, directly reports the surface state of the catalyst at in situ conditions, which at the kinetic phase transition switches from an essentially constantly O-covered (and thus presumably oxidized/reconstructed 35−57) to a mainly CO-covered surface. However, due to the relatively high temperature of our experiment, the CO coverage depends more strongly on the absolute αCO value than the O coverage (for which the higher temperature rather stabilizes any oxide, 55 if formed), since it dictates the equilibrium with CO in the gas phase. 39 Furthermore, it has been demonstrated in a recent combined photoemission electron microscopy (PEEM) and QMS study 58 of the local catalytic ignition during CO oxidation

Figure 5. Correlated ensemble plasmonic nanospectroscopy and mass spectrometry for CO oxidation over Pt. (a) Schematic cross-section of the used array of Au@SiO2−Pt nanostructures. (b) CO2 partial pressure measured by the QMS together with the corresponding spectral position of the plasmonic scattering peak readout, λ, obtained during reaction at 503 K. The data are plotted as a function of the relative CO concentration in the gas flow, αCO = [CO]/([CO] + [O2]), at a constant total reactant concentration of 9% in Ar carrier gas. During the experiment, αCO is swept from 1 to 0 and back to 1, in steps of 0.006 close to αCOcr and in steps of 0.16 otherwise. As the key feature, we observe that a distinct change in trend of the plasmonic nanospectroscopy signal, λ, coincides with the maximum in CO2 partial pressure measured simultaneously by the QMS at αCOcr = 0.04 ± 0.006. (c) The same data as in (b) but plotted as a function of the αCO value for 0 < αCO < 0.5. Both for the αCO up- and down-sweep, the plasmonic peak shift, Δλ, is essentially constant for αCO > αCOcr and then rapidly blue-shifts for αCO < αCOcr to reach a steady-state value beyond αCO ≈ 0.2, where the catalyst activity is essentially zero, as seen from the QMS signal. This is the signature of the kinetic phase transition from a predominantly O-covered to a reduced CO-covered surface.
on low-index Pt surfaces present on a polycrystalline Pt foil that the global QMS CO$_2$ signal obtained from the entire foil (cf. Figure S1 in ref 58) is significantly “smeared” along the CO partial pressure axis (which is equivalent to $\alpha$$_{CO}$ used here) compared to the PEEM intensity measured from single grains with different surface termination ((111), (100), and (110)) within the foil. Using the PEEM intensity together with density functional theory (DFT) calculations the authors also showed that the local kinetic phase diagrams for individual Pt grains in the foil vary significantly depending on their index (cf. Figure 3 in ref 58). Hence, the smearing in the global QMS signal is a consequence of the superposition of the local kinetics of all

Figure 6. Simultaneous single-particle plasmonic nanospectroscopy and ensemble mass spectrometry for CO oxidation over Pt. (a) Schematic cross-section of the used single Au@SiO$_2$–Pt nanostructures. (b–e) Dark-field scattering spectra of the four nanostructures simultaneously under study together with corresponding SEM images (scale bar 50 nm). The scattering spectra were taken for a surface in the CO-covered state (blue) and in the oxidized O-covered state (black). Note the difference in response for the two chemical states. (f–i) Combined plots for scattering peak position $\lambda$ obtained by plasmonic nanospectroscopy (red) for each nanoparticle shown in (b)–(e) and the overall CO$_2$ partial pressure (black) in the reactor measured with the QMS, acquired during a 15 h experiment sweeping $\alpha$$_{CO}$ (blue) from 1 to 0 and then back to 1. We note a different but completely reversible single-nanoparticle response. The black and blue arrows indicate where along the experimental sequence the spectra shown in (b)–(e) were taken.

Figure 7. Simultaneous single-particle plasmonic nanospectroscopy and ensemble mass spectrometry for CO oxidation over Pt at 503 K. (a–d) SEM top view micrographs of the single Au@SiO$_2$–Pt nanostructures used in the experiment depicted in Figure 5. The scale bar is 50 nm. (e–h) CO$_2$ partial pressure measured by the QMS (black) and correlated single-particle plasmonic nanospectroscopy signal (red) as a function of $\alpha$$_{CO}$ up- and down-sweep (upward- and downward-pointing triangles, respectively). The plots are derived based on the raw data depicted in Figure 6 by averaging the QMS signal and the peak shift, $\Delta\lambda$, for each of the four particles over the 15 min long $\alpha$$_{CO}$ steps during the sweep. (i–l) Zoom-in on the kinetic phase transition region ($\alpha$$_{CO}$ = 0–0.5) to highlight the quite different nature of the transition on the different nanoparticles.
grains with a different index present in the foil. In analogy, we argue that the same mechanism is in play in our system since (111), (100), and (110) facets are the dominant surfaces of nanofabricated Pt nanoparticles after annealing, with significant variations in terms of their relative abundance from particle to particle (see Figure 5 in ref 59). Hence, the relatively gradual change of $\Delta \lambda$ across the kinetic phase transition and thus the corresponding gradual change of CO surface coverage measured in our experiment (Figure 5c) are the consequence of the ensemble averaging over ca. 5000 nanoparticles with different facets that exhibit kinetic phase transitions at different $\alpha^{\text{CO}}$.

To enable single-particle plasmonic nanospectroscopy during CO oxidation reaction conditions, we prepared a sample comprising three different areas (Figure S5). In the center is a $4 \times 8$ mm area with the standard nanoparticle surface coverage of ca. 15% obtained by HCL nanofabrication, including a small region close to the edge (ca. 5% of the total surface area) with a particle coverage low enough (i.e., particle–particle distances larger than the diffraction limit) for single-particle plasmonic nanospectroscopy (Figure 6a; see Methods for details of the nanofabrication). To the left and right is a $4 \times 8$ mm area with identical Au@SiO$_2$–Pt nanostructures but at ca. 40% surface coverage to provide enough reaction product for detection by the QMS. In total this means that about 10$^6$ particles are averaged for the QMS readout. Focusing optically on the low-density area of this sample arrangement, we aligned a set of four individual particles within the spectrometer slit, to track their optical response simultaneously and independently (Figure 6b–e). These four particles lie within an area of 35 $\mu$m $\times$ 6 $\mu$m (Figure S5d), which ensures that they experience the same conditions. Using this arrangement, we executed the same experiment as for the ensemble case discussed above, sweeping $\alpha^{\text{CO}}$ (9% total reactant concentration in Ar carrier gas at a constant flow rate of 100 mL/min) in a stepwise fashion in 0.16 and 0.006 $\alpha$-units per step for $\alpha^{\text{CO}} > 0.2$ and $\alpha^{\text{CO}} < 0.2$, respectively, from a CO-rich to a O$_2$-rich condition and back again. The sample temperature was set to 503 K, and we simultaneously monitored the peak position, $\lambda$, of the scattering spectra of the four catalyst nanoparticles and the CO$_2$ reaction product from the corresponding ensemble (Figure 6f–i and Figure S6 for a contour plot of the entire spectral evolution).

To further analyze these raw data and decrease the noise level, we averaged the QMS and $\Delta \lambda$ signals for each of the four particles (Figure 7a–d) for an identical second data set measured at 533 K see Figure S7) over each 15 min long $\alpha^{\text{CO}}$ step and plotted the corresponding averaged $\Delta \lambda$ values as a function of $\alpha^{\text{CO}}$ for both increasing and decreasing $\alpha^{\text{CO}}$ (Figure 7e–h and i–l for zoom-in on a narrower $\alpha^{\text{CO}}$ range). We immediately notice that the overall response looks quite different for the four nanoparticles, which were measured all at the same time. Hence, we can ascribe these differences to their individual response. For example, defining the largest spectral shift between two data points as the kinetic phase transition, we find that it occurs at values ranging from 0.0044 to 0.2 for the $\alpha^{\text{CO}}$ up-sweep on the different particles. At the same time, for the $\alpha^{\text{CO}}$ down-sweep, the difference between particles is smaller. Also hysteresis, however with different width, occurs between the $\alpha^{\text{CO}}$ up- and down-sweeps for all four particles. These observations are in line with our arguments put forward above, that is, that different nanoparticles exhibit variations in terms of the relative abundance of the dominant surface facets, as well as in terms of grain boundaries and related defects (Figure 3). Consequently, they are expected to exhibit the kinetic phase transition at different $\alpha^{\text{CO}}$ in agreement with what we observe in our experiments. Looking even more in detail and comparing the single nanoparticle $\Delta \lambda$ response with the one from the ensemble (cf. Figure 5c), we note that for the single nanoparticles distinct steps in $\Delta \lambda$ and thus in reactant surface coverage occur at the kinetic phase transition, whereas such steps are absent in the ensemble data. Also this observation is thus in line with the main hypothesis of our work, namely, that single-nanoparticle experiments enable insights beyond ensemble averaging, and it demonstrates that the rather smeared out transition in reactant surface coverage reported by $\Delta \lambda$ from the ensemble is the consequence of averaging the rather sharp transitions of the individual nanoparticles, which may occur across a range of $\alpha^{\text{CO}}$, dictated by the particle-specific abundance of grains, defects, and certain surface facets. In this sense, our results for the individual nanoparticles are also well in line with the PEEM study of the polycrystalline Pt foil, for which the kinetic phase transition is rather continuous at the global level and a distinct step at the level of the individual grain with dimensions on the order of 100 $\mu$m (rather than 100 nm as in our case here).58 These observations thus show how the reaction kinetics of simultaneously measured individual model catalyst nanoparticles critically depend on their grain structure and how such information can be obtained at in situ conditions (compared to the to-date used surface science techniques operating at (ultra)high-vacuum conditions such as PEEM58) using plasmonic nanospectroscopy based on a benchtop-type experimental setup comprising an optical microscope and a traditional catalytic flow reactor operating at atmospheric pressure.

**CONCLUSIONS**

In summary, we have presented an experimental setup that combines in situ gas phase single-particle plasmonic nanospectroscopy of individual catalyst nanoparticles with mass spectrometry on a corresponding nanoparticle ensemble in one and the same experiment. In the current design, the system operates at atmospheric pressure and is compatible with temperatures of up to 623 K. We have also developed a nanofabrication method for the crafting of Au@SiO$_2$–Pt hybrid nanostructures to achieve combined sensing function (via an oxide-encapsulated Au nanoantenna) and catalytic function (via a single Pt model catalyst particle on top) within the same structure, to enable plasmonic nanospectroscopy based on the indirect sensing principle at the single-nanoparticle level and in harsh conditions. To demonstrate the capabilities of our setup, nanostructures, and the general experimental approach, we used $\sim 20 \times 70$ nm polycrystalline Pt nanoparticles as the active part in the Au@SiO$_2$–Pt hybrid nanostructure model system, since they offer maximized particle heterogeneity in terms of defects and surface faceting at the single-nanoparticle level. We then applied these structures to investigate the bistable kinetics of the CO oxidation reaction over Pt. As the key results we first found that a characteristic kinetic phase transition can be resolved both at the ensemble and single Pt nanoparticle level as a distinct spectral shift in the plasmonic nanospectroscopy readout and that it occurs at the highest reaction rate identified by the simultaneous QMS readout. This is in agreement with
theory and experimental data obtained using surface science techniques in the corresponding literature. Second, we found that the bistable reaction kinetics and the signature of the kinetic phase transition of simultaneously measured individual model catalyst nanoparticles critically depend on their grain structure, defects, and surface faceting and that the superposition of the individual nanoparticle response induces a significant broadening and smearing in the corresponding kinetic response of a nanoparticle ensemble. This highlights that our experimental approach, which enables studies of catalysts both at the individual nanoparticle level and at the ensemble level in the same system, has the potential to shed light on the role of local catalyst design parameters, such as loading, particle size, shape, and dispersion, as well as metal–support interactions, on reaction kinetics, and how ensemble averaging limits our insights in this respect. To this end, we also briefly summarize the main challenges with the presented approach and developed instrumentation: (i) To maintain the particles of interest in focus during the (typically very long) experiments, in particular at elevated temperature due to thermal expansion effects. Here we predict that the implementation of active focus control solutions will lead to significant improvements in this respect.61 (ii) To obtain enough product from the reaction over these nanofabricated model catalysts to enable online QMS detection. Here the use of microreactors with significantly smaller volumes and thus QMS detection limits could provide interesting solutions.62

In a wider perspective, we predict that our instrument as such, due to the available high temperature and controlled gas environment, and the experimental concept in particular, due to its high sensitivity, will enable single-particle investigations of, for example, the catalyst state during rate oscillations since LSPR is highly sensitive to nanoparticle shape. Furthermore, it will enable studies of the role of the surface state on noble metal catalysts such as Cu and Ag (which themselves are highly plasmonically active) during reaction, of thermally induced nanoalloy formation where LSPR will report on the concurrent change of the complex dielectric function as alloy formation occurs or of the annealing and segregation of alloy components in bimetallic catalysts under reaction conditions via LSPR interface damping effects that occur upon enrichment of a certain element at the surface.63–67

METHODS

Experimental Setup. The main parts of the experimental setup used in this work consisted of a in-house-built transparent flow reactor, a quadrupole mass analyzer (Leda Mass Vacscan), a modified Nikon upright microscope (Eclipse LV150N), and a spectrometer/CCD system (Andor Shamrock 193i spectograph and Andor Newton 920 CCD camera), as shown schematically in Figure 1 and in detail in Figure S1 in the SI. The central part of the flow reactor is a KF40 Duran glass tube (Hositrad), with a round, 25 mm diameter flat borosilicate optical window of 0.2 mm thickness welded in the center to allow collection of images and spectra of individual nanoparticles through the microscope objective in dark-field mode. The reactor is sealed using CF vacuum flanges and Viton O-ring sealedKF flanges, and it operates at atmospheric pressure. Reactants diluted in Ar carrier gas are introduced to the reactor through a gas inlet tube welded onto one of the flanges. The gas flow rate and composition are controlled by a set of mass flow controllers (Bronkhorst Low-AP-flow and EL-flow). The gas outlet is located on the opposite side of the flow reactor to maintain equilibrium plug-flow conditions during the reaction. An in-house-built stainless steel sample holder is mounted inside the reactor tube and equipped with a ceramic resistive heater (Momentive HT01) onto which the sample is clamped. The sample temperature is controlled by a power supply (Instek GW GPS-1850) and a temperature controller (Eurotherm 3216) via a thermocouple-controlled feedback loop to maintain a constant temperature. The latter also prevents defocusing of the sample image during an experiment, which otherwise is induced by thermal expansion and contraction of the sample holder. To enable the mass spectrometric readout, the QMS glass capillary orifice “sniffer” is mounted at the end of the gas transfer line (6.0 mm stainless steel Swagelok tube) to allow only a small and controlled amount of gas leakage into the mass analyzer chamber. The glass capillary orifice was fabricated by inserting the tip of a borosilicate glass tube (2/0.1 mm outer/inner diameter, Hilgenberg) in a propane–oxygen flame in order to shrink the inner diameter to a few micrometers. A detailed description of this method can be found in ref 50. The gas leak rate through the orifice is around 2 × 10^-11 mbar s^-1. With such a leak rate, the pressure inside the QMS analyzer chamber stays below about 6.3 × 10^-6 mbar, which guarantees proper operation of the QMS. Such an orifice design enables efficient molecular flow of the gas from the reactor to the QMS analyzer chamber through the transfer line, and thus a fast response time can be achieved. The gas transfer line outside of the reactor and the QMS analyzer chamber are maintained at 385 K to prevent water condensation.

Mass Spectrometry. The QMS is controlled using a customized LabVIEW™ control program. All QMS raw data were measured on a 200 mbar current of the corresponding gas. Since ≥91% of the gas inside the reactor consists of Ar, the Ar partial pressure (P_Ar) is roughly estimated as being the same as atmospheric pressure (1013.25 mbar). The partial pressure of reactants or product, P, was calculated in mbar by comparison with the Ar signal according to the equation

\[ P = \frac{I}{I_{Ar}} \times 1.27 \times 10^{13} \text{ mbar} \] (1)

where \(I\) and \(I_{Ar}\) are ion currents of gas species i and Ar measured directly by the QMS and S, is the relative sensitivity factor (RSF) of the quadrupole mass analyzer for species i. All shown QMS data are corrected for any background recorded from a control sample without Pt during an identical experimental sequence to the real experiment.

Dark-Field Scattering Spectroscopy. White light from the microscope lamp (Nikon LV-HL50W LL) is used to illuminate the sample, which is mounted inside the reactor close to the optical window, in dark-field mode through the objective (50X Nikon Plano LWD). Scattered light from the sample is collected in a backscattering mode using the objective and then directed toward the entrance slit of the spectograph through a pair of identical 2 in. plano-convex lenses (f = 150 mm, Thorlabs). For single-particle measurements, the sample surface was first imaged at the zero grating position with the spectrometer slit fully open (2500 μm). After subsequently aligning a group of suitable nanoparticles along the center of the slit, the width of the slit was reduced to 450 μm to exclude multiple nanoparticles from being recorded on the same position along the y-axis of the CCD chip. The grating of the spectograph (150 lines/mm, blaze wavelength 800 nm) was then centered at a suitable wavelength around 600–650 nm to acquire a dark-field scattering spectrum from the scatterers on the sample aligned within the slit of the spectograph in the spectral imaging mode. Normalized scattering spectra \(I_s(\lambda)\) from individual particles were thus obtained as a function of wavelength \(\lambda\) using the relation \(I_s(\lambda) = (S - D)/CRS\), where \(S\) is the collected signal from an integrated area with nanoparticle, \(D\) is the signal from the nearby area without nanoparticle (dark signal for background correction taken from an area with identical pixel width but without particles), and CRS is the signal collected from the diffuse white certified reflectance standard bright reference sample (Labsphere SRS-99-0200). CRS is used in order to correct the signal for the lamp spectrum. The acquisition time for each spectrum was 10 to 15 s depending on the brightness of the particle. The obtained single-particle scattering spectra were fitted with a Lorentzian function (±75 nm from the peak position) to derive information about the temporal evolution of the peak position.68

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For scattering measurements from nanoparticle ensembles, the slit width of the spectrometer was reduced to 100 μm to prevent saturation of the CCD sensor chip. To calculate normalized scattering spectra $I_s$, the background signal $D$ was collected from a blank substrate without any particles, under the same conditions as when the signal $S$ was collected. The same CRS as in single-particle measurements was used for spectrum correction. Peak position was extracted using the same fitting procedure as in single-particle measurements.

**Sample Nanofabrication.** The Au@SiO$_2$–Pt nanostructures were fabricated on a thermally oxidized silicon substrate (100 nm oxide thickness) using a tailored variant of the hole–mask colloidal lithography (HCL) method that is described in detail elsewhere. The main new step in the present nanofabrication route is the growth of a 10 nm thin SiO$_2$ layer through the nanofabrication mask generated using the standard HCL process, encapsulating a plasmonic nanoantenna particle grown before. The key step enabling this through-mask sputtering without compromising the final lift-off step is a prolonged oxygen plasma etch ($90$ s, $50$ W, $250$ mTorr, Plasma-Rem 700 (Microresist Technology GmbH), isopropyl alcohol, and acetone). It is important to note that the single-particle region of 5% Rem 700 (Microresist Technology GmbH), isopropyl alcohol, and acetone was formed due to faster drying at the edge when blow-drying the Pt catalyst. Finally, lift-off was carried out at corresponding temperatures. The CO$_2$ QMS signal from the sample with Pt from the control sample without Pt catalyst.

**CO Oxidation Experiments on Pt Catalyst Nanoparticle Ensemble.** Prior to sweeping $\alpha^{CO}$ in the kinetic phase transition experiments, the sample was exposed to 20 cycles of alternating 9% CO ($6.0$ purity in Ar) and 9% O$_2$ ($6.0$ purity in Ar) pulses of 15 min duration, followed by a 1 h pulse of 9% CO$_2$, in order to activate the catalyst and reach a stable optical signal. During this treatment, the sample temperature was kept at $553$ K. For the $\alpha^{CO}$ sweep, the temperature was kept constant at the set value, and the sample was exposed to a constant total reactant concentration [CO + O$_2$] of 9%, at a constant flow rate of 100 mL/min. A scattering spectrum was simultaneously collected every 60 s using an integration time of 0.1 s and 10 accumulations for the CCD. With the QMS, the ion currents of CO, CO$_2$, air/N$_2$, O$_2$, and Ar inside the reactor were continuously measured at a mass value of 28, 44, 14, 32 and 40, respectively, with a time resolution of 5 s. The QMS was operated in SEM mode. To subtract the CO$_2$ background signal, measurements on a control sample without Pt under the exact experiment conditions were also carried out at corresponding temperatures. The CO$_2$ QMS signal shown in this article for both ensemble and single-particle experiments was corrected by subtracting the CO$_2$ signal from the sample with Pt from the control sample without Pt catalyst.

**CO Oxidation Experiments on Single Pt Catalyst Nanoparticles.** Nanoparticles in the sample region of low particle density as shown in Figure 3d were chosen to achieve well-separated (>20 pixels) diffraction-limited spots on the CCD sensor chip, using the 50X objective of the microscope. The sample stage was adjusted to align the image of the chosen nanoparticles within the view of the spectrometer slit set to a 450 μm opening. The scattered light from each nanoparticle dispersed by the grating was collected in spectral image mode with 14 s integration time and accumulated with 4 acquisitions every 60 s. The same conditions as in the ensemble CO oxidation experiment were applied to compare the two results directly. A normalized scattering spectrum from each nanoparticle was extracted and fitted according to the procedure described in the previous dark-field scattering spectroscopy method section to extract the peak position.

**References**


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**Notes**

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