



A simplicity-guided cocktail approach toward multicolor fluorescent systems

Downloaded from: <https://research.chalmers.se>, 2024-04-19 18:00 UTC


Citation for the original published paper (version of record):

Naren, G., Li, S., Andreasson, J. (2020). A simplicity-guided cocktail approach toward multicolor fluorescent systems. *Chemical Communications*, 56(23): 3377-3380.
<http://dx.doi.org/10.1039/c9cc10040g>

N.B. When citing this work, cite the original published paper.



A simplicity-guided cocktail approach toward multicolor fluorescent systems†

Gaowa Naren, Shiming Li and Joakim Andréasson *

Cite this: *Chem. Commun.*, 2020, **56**, 3377

Received 28th December 2019,
Accepted 13th February 2020

DOI: 10.1039/c9cc10040g

rsc.li/chemcomm

A molecular cocktail containing two photochromic diarylethene derivatives that displays multicolor emission spanning blue-green to orange in a color-correlated fashion has been devised. The function does not rely on excited state communication such as energy transfer reactions, which is the typical case for similar systems. Instead, harnessing the intrinsic fluorescent and photochromic properties of the two individual diarylethene derivatives run in parallel is enough to realize the color changes. This offers an extremely flexible situation as for the choice of the fluorophores and their respective concentrations. The function is conveniently demonstrated in bulk solution at μM concentrations, where a single light source serves as the color changing stimulus.

Over the last decade, there has been an increasing interest in molecular systems displaying stimuli-responsive changes in the emission color. Main drivers for this development are potential applications in optical data storage, bioimaging, super-resolution microscopy, sensors, anticounterfeiting, and more.^{1–11} Compared to binary on-off (single-color) fluorescence switching, stimuli-responsive multicolor systems offer many upsides such as allowing for ratiometric methods,^{12–15} full-color reproduction,^{16–18} and the possibility to introduce a high degree of unpredictability in the associated spectral changes, a highly sought-after property in data-security contexts.^{1,19,20}

For multicolor fluorescence photoswitching, the color changes can be brought about in two principally different ways: color-specific and color-correlated switching.⁴ The latter allows for a much more distinct color change, as the increase in emission in one wavelength region is concerted, or correlated, with the disappearance of the emission in another wavelength region. A common strategy for color-correlated systems is to make use of FRET processes,^{21–24} typically by employing a “static”

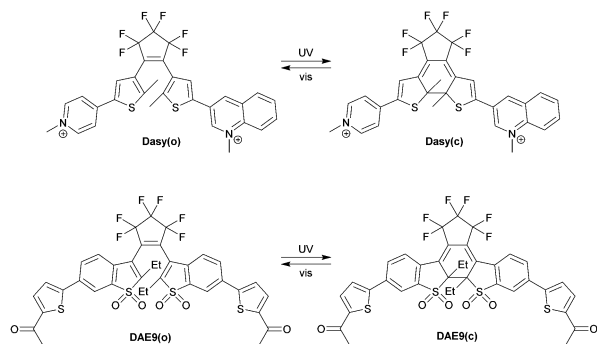
fluorescent dye as the energy donor and a molecular photoswitch (photochromic molecule) as the corresponding acceptor. Isomerization of the photochromic acceptor between the colorless passive form and the colored, fluorescent FRET-active form allows in principle for convenient toggling of the emission between 100% donor-centered to 100% acceptor-centered. Efficient FRET reactions, however, require the donor and the acceptor in close proximity, typically at around 40 Å or closer. Several supramolecular and covalently linked constructs have been presented,^{16,25–31} the common denominator being a careful organization of the system into a suitable geometry that allows for efficient FRET reactions to occur. Most often these approaches imply complicated synthetic procedures and the loss of system flexibility. Here we report a novel, extremely facile means of color-correlated photoswitching by simply mixing a cocktail of two photochromic diarylethene (DAE) derivatives³² in acetonitrile bulk solution at μM concentration. We harness merely the intrinsic fluorescent and photochromic properties of the two photoswitches without the need to fine-tune excited state reactions such as the above-mentioned FRET processes. In addition, the system is controlled by a single light source at 381 nm triggering the color change, and a second light source at 410 nm for emission readout with virtually non-destructive character.

The structures and the isomerization scheme of the two DAE derivatives used in this study are shown in Scheme 1. Herein, we refer to the compounds as **Dasy** (DAE asymmetric) and **DAE9** (adopted from the nomenclature in the original report).³³ Both exist in an open isomeric form absorbing almost exclusively in the UV region (see Fig. 1 for spectral properties). UV exposure triggers isomerization from the open to the closed form, and the reverse reaction is induced by visible light. **Dasy** is fluorescent in the open isomeric form **Dasy(o)** with a quantum yield of 0.11 whereas the closed ditto **Dasy(c)** displays no detectable emission. The opposite is true for **DAE9** where there is no detectable emission from **DAE9(o)** but **DAE9(c)** emits fluorescence with a quantum yield of 0.37. This value is significantly lower compared to the fluorescence quantum yield in 1,4-dioxane (0.61).³³ Similar DAE derivatives have been reported to show the same trend, that is,

Department of Chemistry and Chemical Engineering, Chemistry and Biochemistry, Chalmers University of Technology, SE-41296 Göteborg, Sweden.
E-mail: a-son@chalmers.se

† Electronic supplementary information (ESI) available: Synthesis, experimental section, and supporting figures. See DOI: 10.1039/c9cc10040g





Scheme 1 Structures and isomerization scheme of the studied compounds.

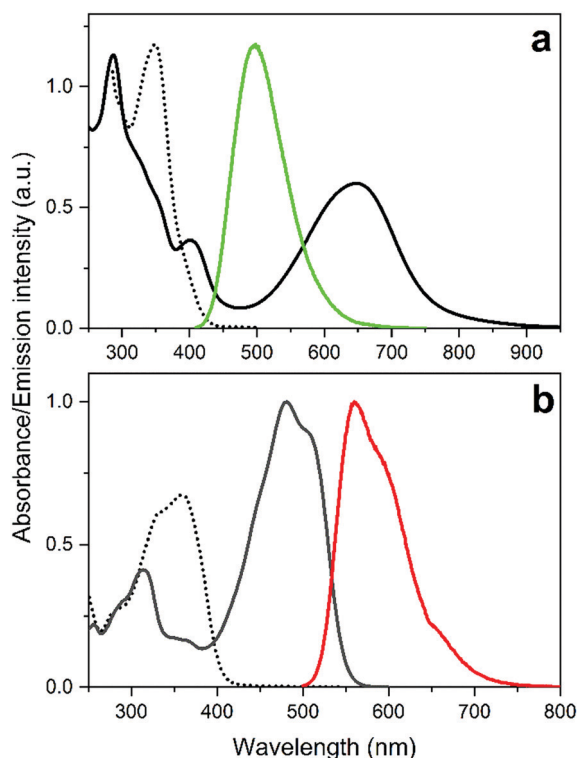


Fig. 1 Absorption and emission spectra of **Dasy** (a) and **DAE9** (b). Absorption spectra of the open isomers (dotted black lines) and the closed isomers (solid black lines) are shown together with the emission spectra (green solid line for **Dasy(o)** and red solid line for **DAE9(c)**).

a decrease in the fluorescence quantum yield with increasing solvent polarity.³⁴

As for the emission colors, **Dasy(o)** emits blue-green light centered at around 497 nm, and **DAE9(c)** emits in the orange region with emission maximum at 561 nm. The CIE coordinates of the respective spectra are shown in Fig. S12, ESI†. We would like to emphasize that a fluorescence quantum yield of 0.11 is unusually high for the open isomer of a regular perfluoro DAE derivative. In fact, we know of only two reported examples with a higher number.^{35,36} These compounds, however, suffer from a UV-induced photostationary state poorly enriched in the closed form, implying that “on-off” switching of the emission is not possible.

Exposing a cocktail solution of **Dasy(o)** and **DAE9(o)** to UV light would imply initial emission of blue-green fluorescence from **Dasy(o)** only. This emission will gradually disappear, as UV exposure triggers also the isomerization to the non-fluorescent **Dasy(c)** isomer. Concomitant to the decrease in blue-green emission intensity, an increase in the orange emission from **DAE9(c)** is expected, due to the UV-induced formation of this species. Extended UV irradiation would eventually yield a total disappearance of the blue-green emission (given that the photostationary distribution of **Dasy** is 100% shifted to the non-fluorescent **Dasy(c)** isomer) and exclusively intense orange emission from **DAE9(c)**. Please note that this color change is expected to occur without the need for FRET reactions. It is simply a matter of balancing the rates of the two isomerization reactions, **Dasy(o)** → **Dasy(c)** and **DAE9(o)** → **DAE9(c)** allowing for a continuous color change. The choice of irradiation wavelength indeed allows for tuning the relative rates of these isomerization reactions (Fig. S13, ESI† and text below).

As for the reading process, that is, excitation for emission readout, a constant overall emission intensity throughout the process is desirable. Moreover, reading the sample in a non-destructive fashion, implying that the excitation light for emission readout does not induce changes in the isomeric distribution, would also add to the appeal. The former matter reduces to choosing the excitation wavelength as to match the brightness of the two fluorophores, whereas the latter implies exciting the cocktail at a wavelength with minimal absorption of the photochromically active species. We have detailed in the ESI† how the photophysical properties of the photoswitches are used as a guide to the experimental conditions for optimal performance according to above.

Fig. 2 shows the evolution of the overall emission spectra of a sample containing *ca.* 35 μM **Dasy** and 2.3 μM **DAE9** with increasing exposure time to 381 nm UV light. Excitation for emission readout was done at 410 nm. It is clearly seen that the initial spectrum (green) contains no or only minor contribution from **DAE9(c)** but is instead dominated by **Dasy(o)**, whereas the final spectrum (red) shows fluorescence from almost exclusively **DAE9(c)**. Moreover, the overall emission intensities are comparable at all times as shown in Fig. S14, ESI† and the color change occurs in a continuous fashion from blue-green to orange, as is evident from Fig. S15, ESI†. This clearly shows that the rates of the isomerization reactions **Dasy(o)** → **Dasy(c)** and **DAE9(o)** → **DAE9(c)** occur on a similar timescale, which is also seen by comparing the individual isomerization rates from UV-vis absorption measurements where the rates differ by only a factor 2 (see Fig. S13, ESI†). The isomerization quantum yields for the closing reactions are 0.52 and 0.13 for **Dasy** and **DAE9**, respectively. Resetting the sample to the initial state containing only **Dasy(o)** and **DAE9(o)** is conveniently done by exposure to visible light, in this case at 523 nm, allowing for repeated operation with excellent reproducibility (Fig. S16 and S17, ESI†). The sample can be exposed to at least five switching cycles without substantial photodegradation (Fig. S18, ESI†).

The stability of the system with respect to color changes over time both in the absence and in the presence of the excitation



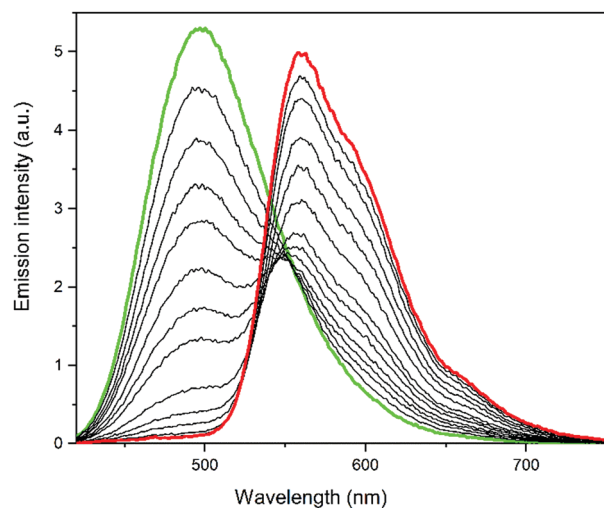


Fig. 2 Time evolution of the emission spectra for the **Dasy** + **DAE9** cocktail in acetonitrile under exposure to 381 nm UV light for isomerization purpose. All spectra were recorded upon excitation at 410 nm after exposure to 381 nm light for 0 s (green line), 5 s, 10 s, 15 s, 20 s, 30 s, 40 s, 50 s, 80 s, 110 s, 140 s, 200 s, 260 s, and 350 s (red line).

light for emission readout was investigated (Fig. S19–S23, ESI†). The former reflects the thermal stability of the system when the photoswitches are in the closed isomeric form. Here, samples of **DAE9(o)** and **Dasy(o)** were converted to the photostationary state using UV light at 365 nm (both photoswitches display virtually quantitative conversion to the closed isomers) and left for 24 hours in the dark. No detectable changes in the absorption spectra were observed for **DAE9**, whereas **Dasy** experienced *ca.* 10% decrease in the absorption of the closed isomer. The absorbance in the wavelength region 330–380 nm where **Dasy(o)** displays its characteristic strong absorption band also decreased, strongly signalling that the spectral changes are due to the irreversible formation of a by-product.³⁷ The system displays excellent color stability with respect to excitation for emission readout, that is, only minor changes in the emission spectra were observed after recording up to 80 spectra with excitation at 410 nm. There is, of course, an inevitable trade-off between absolute emission intensity and color stability, and in this study the excitation wavelength was chosen as to optimize the color stability without sacrificing too much in the signal-to-noise of the fluorescence signal.

In order to demonstrate that the observed color changes do not rely on FRET reactions of any kind, time-resolved single photon counting experiments were undertaken (see Fig. S24 and S25, ESI†). The fluorescent lifetimes of the monomers alone in acetonitrile are 1.8 ns for **DAE9(c)** and 5.4 ns for **Dasy(o)**. The lifetimes for **DAE9(c)** in the cocktail together with **Dasy(o)** and **Dasy(c)** are both 1.8 ns, whereas **Dasy(o)** has lifetimes of 5.8 ns and 5.7 ns in the cocktail with **DAE9(o)** and **DAE9(c)**, respectively. Thus, it is clear that quenching by FRET reactions do not occur.

Finally, the choice of photoswitches should be commented on. The first prerequisite is that one of the two must be emissive in the colorless form, whereas the other must be emissive in the

colored form. Indeed, a photoswitch that displays intense emission in both isomeric forms would be usable too, but to date, there are no photochromic compounds with this favourable property. Moreover, it would imply a loss in the overall flexibility as the total concentration of the two emissive species would remain constant through the isomerization process. Using **Dasy** and **DAE9** demonstrates the usefulness of the cocktail approach to finetune both the relative concentrations and the excitation wavelengths to compensate for differences in the emission- and the isomerization quantum yields of the ingredient photoswitches. Please note that although **Dasy(o)** used herein displays unusually high fluorescence quantum yield compared to the typical open form of any DAE derivative, an equivalent function would be achievable also for less fluorescent compounds, but then at a higher concentration to compensate. Alternatively, a covalently linked dyad comprising a fluorophore covalently linked to a non-fluorescent photochromic compound to constitute a photoswitchable FRET (or photoinduced electron transfer) donor–acceptor pair could be used. Note, however, that this approach would impose additional constraints to the design criteria, as the excited state interactions between the donor and the acceptor would have to be carefully tuned.

Moreover, the structural properties of **Dasy** and **DAE9** differ, in particular with respect to **Dasy** being a dication whereas **DAE9** is charge neutral. This presents challenges as for many approaches to self-assemble both compounds in confined volumes, which is a requirement for efficient FRET reactions to occur. The herein described approach overcomes also this limitation. It should be noted, however, that the ease of operation requires a medium in which the two photoswitches are fairly evenly distributed, as regions of high local concentration of both chromophores will interfere with the overall function by the onset of excited state reactions such as FRET and photoinduced electron transfer.

In summary, we have presented a novel approach to multi-color fluorescent photoswitching in a color-correlated fashion. The function does not rely on FRET reactions to occur, which allows for an extremely simple experimental protocol where two fluorescent DAE derivatives are mixed in bulk acetonitrile solution at μM concentrations. Hence, tedious synthesis of covalently linked constructs or the formation of supramolecular systems are avoided. A single light source is responsible for the isomerization-induced color changes and the emission readout process is virtually non-destructive. The all-photonic nature of the external stimulus presents the typical advantages (waste-free, non-invasive, and remote operation together with the unsurpassed spatiotemporal control) compared to systems that are driven by chemical or mechanical triggers.

JA acknowledges financial support from the Swedish Research Council VR (Grant # 2016-0360). The authors thank Dr Tamara Pace and Dr Vilhelm Müller for participating in the initial studies of **Dasy**.

Conflicts of interest

There are no conflicts to declare.



Notes and references

‡ There are examples of other DAE derivatives with similar or higher fluorescence quantum yields in the open isomeric form.^{32,38–44} In these compounds, however, the emission is originating from a distinct fluorophore associated with the photoswitch, and not integral part of the same.

- 1 J. Andréasson and U. Pischel, *Chem. Soc. Rev.*, 2018, **47**, 2266–2279.
- 2 N. Hildebrandt, C. M. Spillmann, W. R. Algar, T. Pons, M. H. Stewart, E. Oh, K. Susumu, S. A. Diaz, J. B. Delehanty and I. L. Medintz, *Chem. Rev.*, 2017, **117**, 536–711.
- 3 X. S. Hou, C. F. Ke, C. J. Bruns, P. R. McGonigal, R. B. Pettman and J. F. Stoddart, *Nat. Commun.*, 2015, **6**, 6884.
- 4 D. Kim and S. Y. Park, *Adv. Opt. Mater.*, 2018, **6**, 1800678.
- 5 C. H. Li and S. Y. Liu, *Chem. Commun.*, 2012, **48**, 3262–3278.
- 6 V. K. Praveen, C. Ranjith and N. Armaroli, *Angew. Chem., Int. Ed.*, 2014, **53**, 365–368.
- 7 S. Silvi, M. Baroncini, M. La Rosa and A. Credi, *Top. Curr. Chem.*, 2016, **374**, 65.
- 8 Z. Y. Tian, W. W. Wu and A. D. Q. Li, *ChemPhysChem*, 2009, **10**, 2577–2591.
- 9 Y. L. Zhang, K. Q. Zhang, J. Wang, Z. Y. Tian and A. D. Q. Li, *Nanoscale*, 2015, **7**, 19342–19357.
- 10 D. Kim, K. Jeong, J. E. Kwon, H. Park, S. Lee, S. Kim and S. Y. Park, *Nat. Commun.*, 2019, **10**, 3089.
- 11 Z. Xu, D. Gonzalez-Abradelo, J. Li, C. A. Strassert, B. J. Ravoo and D. S. Guo, *Mater. Chem. Front.*, 2017, **1**, 1847–1852.
- 12 W. Zhou, J. Zhu, D. Fan, Y. Teng, X. Zhu and S. Dong, *Adv. Funct. Mater.*, 2017, **27**, 1704092.
- 13 L. Hou, Y. Qin, J. Li, S. Qin, Y. Huang, T. Lin, L. Guo, F. Ye and S. Zhao, *Biosens. Bioelectron.*, 2019, **143**, 111605.
- 14 S. A. Diaz, L. Giordano, T. M. Jovin and E. A. Jares-Erijman, *Nano Lett.*, 2012, **12**, 3537–3544.
- 15 S. A. Diaz, L. Giordano, J. C. Azcarate, T. M. Jovin and E. A. Jares-Erijman, *J. Am. Chem. Soc.*, 2013, **135**, 3208–3217.
- 16 M. Yu, P. Zhang, B. P. Krishnan, H. Wang, Y. Gao, S. Chen, R. Zeng, J. Cui and J. Chen, *Adv. Funct. Mater.*, 2018, **28**, 1804759.
- 17 G. Naren, C. W. Hsu, S. M. Li, M. Morimoto, S. C. Tang, J. Hernando, G. Guirado, M. Irie, F. M. Raymo, H. Sundén and J. Andréasson, *Nat. Commun.*, 2019, **10**, 3996.
- 18 J. X. Bu, K. Watanabe, H. Hayasaka and K. Akagi, *Nat. Commun.*, 2014, **5**, 3799.
- 19 H. Wu, Y. Chen, X. Dai, P. Li, J. F. Stoddart and Y. Liu, *J. Am. Chem. Soc.*, 2019, **141**, 6583–6591.
- 20 C. W. Hsu, C. Sauvee, H. Sundén and J. Andréasson, *Chem. Sci.*, 2018, **9**, 8019–8023.
- 21 M. Bälter, S. M. Li, M. Morimoto, S. C. Tang, J. Hernando, G. Guirado, M. Irie, F. M. Raymo and J. Andréasson, *Chem. Sci.*, 2016, **7**, 5867–5871.
- 22 T. Wu, G. Zou, J. M. Hu and S. Y. Liu, *Chem. Mater.*, 2009, **21**, 3788–3798.
- 23 L. Y. Zhu, W. W. Wu, M. Q. Zhu, J. J. Han, J. K. Hurst and A. D. Q. Li, *J. Am. Chem. Soc.*, 2007, **129**, 3524–3526.
- 24 S. H. Lee, H. T. Bui, T. P. Vales, S. Cho and H. J. Kim, *Dyes Pigm.*, 2017, **145**, 216–221.
- 25 H. Wang, P. Zhang, B. P. Krishnan, M. Yu, J. Liu, M. Xue, S. Chen, R. Zeng, J. Cui and J. Chen, *J. Mater. Chem. C*, 2018, **6**, 9897–9902.
- 26 S. Wang, T. Li, X. Zhang, L. Ma, C. Li, X. Yao, D. Cao and X. Ma, *ChemPhotoChem*, 2019, **3**, 568–574.
- 27 Z. Lin, H. Wang, M. Yu, X. Guo, C. Zhang, H. Deng, P. Zhang, S. Chen, R. Zeng, J. Cui and J. Chen, *J. Mater. Chem. C*, 2019, **7**, 11515–11521.
- 28 P. Deng, J. Sun, J. Chen, X. Zou and L. Liao, *Carbohydr. Polym.*, 2019, **210**, 379–388.
- 29 D. Kim, J. E. Kwon and S. Y. Park, *Adv. Funct. Mater.*, 2018, **28**, 1706213.
- 30 S. A. Diaz, F. Gillanders, K. Susumu, E. Oh, I. L. Medintz and T. M. Jovin, *Chem. – Eur. J.*, 2017, **23**, 263–267.
- 31 S. Ishida, T. Fukaminato, D. Kitagawa, S. Kobatake, S. Kim, T. Ogata and S. Kurihara, *Chem. Commun.*, 2017, **53**, 8268–8271.
- 32 M. Irie, T. Fukaminato, K. Matsuda and S. Kobatake, *Chem. Rev.*, 2014, **114**, 12174–12277.
- 33 K. Uno, H. Niikura, M. Morimoto, Y. Ishibashi, H. Miyasaka and M. Irie, *J. Am. Chem. Soc.*, 2011, **133**, 13558–13564.
- 34 M. Morimoto, Y. Takagi, K. Hioki, T. Nagasaka, H. Sotome, S. Ito, H. Miyasaka and M. Irie, *Dyes Pigm.*, 2018, **153**, 144–149.
- 35 R. Wang, S. Pu, G. Liu, S. Cui and W. Liu, *J. Photochem. Photobiol., A*, 2012, **243**, 47–55.
- 36 K. Shibata, L. Kuroki, T. Fukaminato and M. Irie, *Chem. Lett.*, 2008, **37**, 832–833.
- 37 M. Irie, T. Lifka, K. Uchida, S. Kobatake and Y. Shindo, *Chem. Commun.*, 1999, 747–750.
- 38 C. C. Ko, W. H. Lam and V. Wing-Wah Yam, *Chem. Commun.*, 2008, 5203–5205.
- 39 R. Kanazawa, T. Nakashima and T. Kawai, *J. Phys. Chem. A*, 2017, **121**, 1638–1646.
- 40 S. Huang, Z. Li, S. Li, J. Yin and S. Liu, *Dyes Pigm.*, 2012, **92**, 961–966.
- 41 M. T. W. Milder, J. L. Herek, J. Areephong, B. L. Feringa and W. R. Browne, *J. Phys. Chem. A*, 2009, **113**, 7717–7724.
- 42 C. T. Poon, W. H. Lam, H. L. Wong and V. W. W. Yam, *J. Am. Chem. Soc.*, 2010, **132**, 13992–13993.
- 43 H. H. Liu and Y. Chen, *J. Phys. Chem. A*, 2009, **113**, 5550–5553.
- 44 T. Fukaminato, S. Ishida and R. Metivier, *NPG Asia Mater.*, 2018, **10**, 859–881.

