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The role of connectivity in significant bandgap narrowing for fused-pyrene based non-fullerene acceptors toward high-efficiency organic solar cells†

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Great attention has been paid to developing low bandgap non-fullerene acceptors (NFAs) for matching wide bandgap donor polymers to increase the photocurrent and therefore the power conversion efficiencies (PCEs) of NFA organic solar cells, while pyrene-core based acceptor–donor–acceptor (A–D–A) NFAs have been mainly reported via the 2,9-position connection due to their bithieno[3',2'-b']thienyl[a,h]pyrene fused via a five-membered ring bridge at the ortho-position of pyrene as the representative one named FPIC5, which has prohibited further narrowing their energy gap. Herein, an acceptor FPIC6 was exploited by creating the 1,8-position connection through fusing as bithieno[3',2'-b']thienyl[f-g,m-n]pyrene linked at the bay-position via a six-membered bridge, with enhanced push–pull characteristics within such A–D–A structure. As a structural isomer of FPIC5, FPIC6 exhibited a much lower bandgap of 1.42 eV (1.63 eV for FPIC5). Therefore, the photocurrent and PCE of PTB7-Th:FPIC6 cells were improved to 21.50 mA cm⁻² and 11.55%, respectively, due to the balanced mobilities, better photoluminescence quenching efficiency and optimized morphology, which are both ~40% better than those of PTB7-Th:FPIC5 cells. Our results clearly proved that a pyrene fused core with 1,8-position connection with electron-withdrawing end groups instead of 2,9-position connection is an efficient molecular design strategy to narrow the optical bandgap and improve the photovoltaic performance of NFA based OSCs.

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1. Introduction

A breakthrough milestone for boosting the performance of organic solar cells (OSCs) in 2015 was the development of a non-fullerene acceptor (NFA) abbreviated as ITIC by Lin and Zhan *et al.* and the follow-on derivatives with their concept of

fused-ring electron acceptors (FREAs).^{1–3} Such a breakthrough is based on not only the significantly improved complementary optical absorption matches between the donor and acceptor materials and the finely tunable LUMO levels of NFAs, but also their high electron mobility.^{4–17} Therefore such a perfect combination of wide bandgap polymers as the donor

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and low bandgap NFAs has pushed the efficiency of OSCs up to 18%.^{18–21}

Among the state-of-the-art donor-core moieties for FREAs such as indacenodithiophene (IDT), indacenodithieno[3,2-*b*]thiophene (IDTT), oligomeric IDT, and their structural analogues, pyrene, as a cheap and commercially available polycyclic aromatic hydrocarbon composed of four fused benzene rings, has gained great attention in the rational molecular design of organic semiconductors due to its good structural planarity, excellent thermal stability, largely decreased ionization potential, extensive electron delocalization and fast charge-transport ability.^{22–28} As a result, pyrene derivatives usually exhibit strong π - π stacking, high degree of crystallinity, and enhanced charge carrier mobilities, and therefore have been widely used in organic field effect transistors (OFETs), organic light emitting diodes (OLEDs), and hole-transporting materials for perovskite solar cells and OSCs.^{27–31} This gave birth to a medium-bandgap FPIC (termed as **FPIC5** in this paper, Scheme 1) by Lian and Zhan *et al.* in 2018, which was based on pyrene fused with two thieno[3,2-*b*]thiophenes *via* five-member-ring bridges of di(*para*-hexyl)phenylcyclopentadienyl at the two ortho-positions of pyrene[*a,h*] (2,9-positions).³² The optical bandgap of **FPIC5** was found to be medium at 1.63 eV, leading to a low short-circuit current density (J_{SC}) of 15.3 mA cm⁻² with a moderate PCE of 8.45% for solar cells based on **PTB7-Th:FPIC5**. Another two NFAs with the pyrene core connecting with end groups *via* 2,9-positions named Py-1 and Py-2 also exhibited a wide bandgap of 1.95 and 1.97 eV respectively, leading to low J_{SC} and PCE.³³ On the other hand, “DTP-IC-4Ph” (dithieno[*f,g,m-n*](di(*para*-hexyl)phenyl)cyclohexapyrene-IC) from Guo *et al.*³⁴ with the pyrene

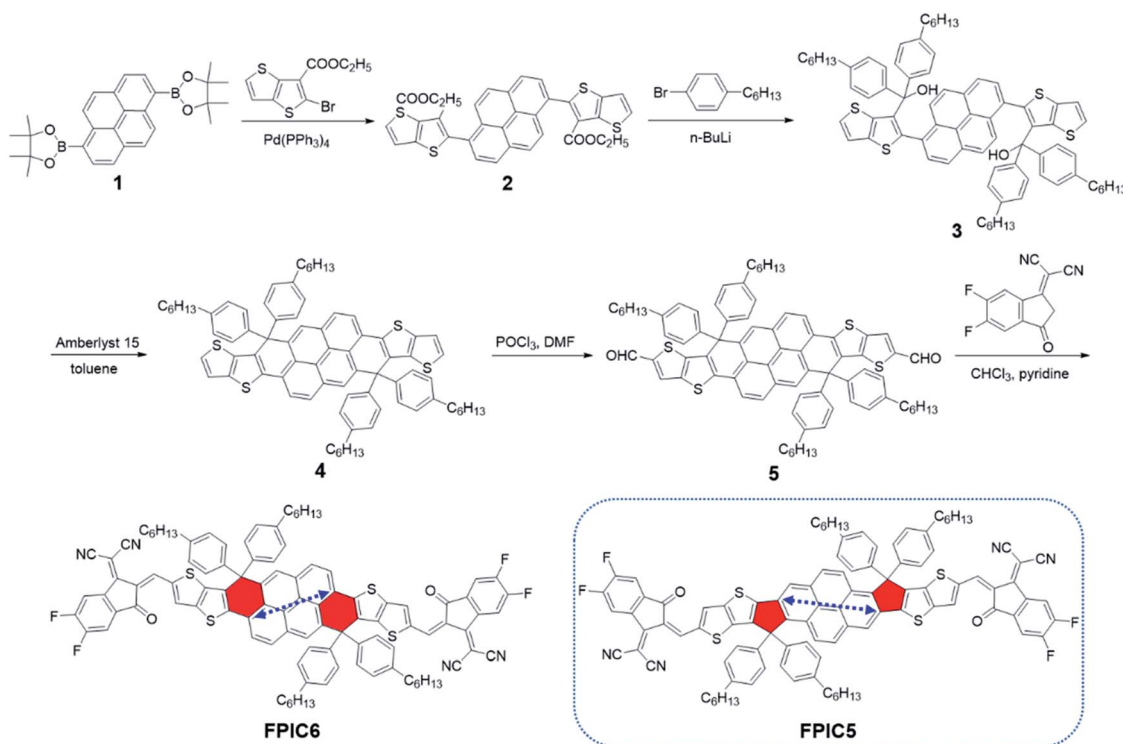
core connecting with end groups *via* 1,8-positions presented a notably red-shifted absorption spectrum and thus a clearly improved J_{SC} of 17.32 mA cm⁻² and PCE of 10.37%. However, the reasons for the red-shifted spectra and how the connection position of the pyrene core can influence the properties of the resulting molecules have not been systematically studied.

Above all, isomerization of pyrene core based NFAs with connection varying from 2,9-positions to 1,8-positions could dramatically enhance both conjugation and connectivity, leading to low-bandgap conjugated molecules. This inspired our design idea on developing a new small molecule acceptor **FPIC6** in this work, using pyrene condensed with two thieno[3,2-*b*]thiophenes *via* six-membered ring bridges as the central core and 2-(5,6-difluoro-3-oxo-2,3-dihydro-1*H*-inden-1-ylidene) malononitrile (2FIC) as the end-groups through 1,8 sites (Scheme 1). It is noteworthy that **FPIC6** with its 1,8-position connection exhibited an \sim 119 nm red-shifted absorption when compared to the reported isomer **FPIC5** with 2,9-position connection, leading to a narrow bandgap of 1.42 eV (obtained from optical absorption) and 1.53 eV (determined by CV). As a result, simple binary **PTB7-Th:FPIC6** solar cells exhibited a much higher photocurrent (21.5 vs. 15.3 mA cm⁻²) and PCE (11.55% vs. 8.45%) than its structural isomer **FPIC5**.

2. Results and discussion

2.1. Materials synthesis and characterization

The synthetic route to **FPIC6** is shown in Scheme 1, and the experimental details are presented in the Experimental section and ESI.† A Suzuki coupling reaction between 1,6-bis(4,4,5,5-



Scheme 1 Synthetic route to small molecule acceptor **FPIC6** and the chemical structure of **FPIC5**.



tetramethyl-1,3,2-dioxaborolan-2-yl)pyrene (**1**)³⁵ and ethyl 2-bromothiopheno[3,2-*b*]thiophene-3-carboxylate could give diethyl 2,2'-(pyrene-1,6-diyl)bis(thiopheno[3,2-*b*]thiophene-3-carboxylate) (**2**) with a yield of 88.7%. Compound **2** was reacted with 4-hexylphenyl lithium to give the tertiary alcohol **3**, which could be converted to decacyclic compound **4** *via* the Amberlyst 15 assisted Friedel-Crafts alkylation reaction. Then, the dialdehyde compound **5** was prepared by means of a Vilsmeier-Haack reaction of compound **4** with phosphorus oxychloride (POCl₃) and *N,N*-dimethylformamide (DMF) with a yield of 84.1%. A Knoevenagel condensation reaction between compound **5** and 2FIC afforded the final product **FPIC6** (yield: 77.0%). The chemical structures of intermediates and **FPIC6**, and their purities were confirmed by NMR and mass spectra. **FPIC6** is readily soluble in common organic solvents such as dichloromethane, chloroform and chlorobenzene. As shown in Fig. S1,[†] the decomposition temperature (T_d , 5% wt loss) of **FPIC6** is 377 °C, indicating its excellent thermal stability for photovoltaic applications.

2.2. Optical properties

For the UV-vis-NIR absorption spectra of **FPIC6** both in dilute chloroform solution and as thin films (as displayed in Fig. 1a), compared with the weak absorption for the π - π^* transition below 500 nm, there is a clear main peak at 723 nm for its solution arising from the intramolecular charge transfer (ICT) from the electron-donating pyrene core to the terminal electron accepting moieties. As shown in Fig. S2,[†] the temperature-dependent UV-vis absorption of **FPIC6** in solution discloses slight blue-shifts ranging from 2 to 10 nm with increasing temperature from 20 °C to 100 °C, indicating that there is no evidence of even moderate molecular aggregation in the dilute solutions, and all the main peaks originated from the ICT as stated above. When seeing the very much alike absorption spectrum of its counterpart **FPIC5** in the solution,³² the main peak is evidently red-shifted from 645 nm for **FPIC5** to 723 nm for **FPIC6** (with a molar extinction coefficient of $1.73 \times 10^5 \text{ M}^{-1} \text{ cm}^{-1}$). This reveals both stronger electron push-pull effects and more extended conjugation through improved

connectivity (will be discussed more later on) and therefore an efficiently narrowed energy bandgap for **FPIC6** compared to the five-membered ring fused **FPIC5**. Meanwhile, the **FPIC6** film exhibits strong absorption in the region of 600–870 nm with an absorption peak at 785 nm. From solution to thin films, **FPIC6** exhibits a large bathochromic shift of 62 nm, demonstrating strong intermolecular interaction in its solid state.^{36–44} It is noteworthy that the absorption peak is significantly 119 nm red-shifted, compared with its isomer **FPIC5** (Table 1). This reflects that **FPIC6** with the six-membered ring bridge possesses a more effective intramolecular charge transfer effect compared to five-membered ring bridged **FPIC5**.³² The optical bandgap (E_g^{opt}) of **FPIC6** was calculated to be 1.42 eV from the absorption edge of 871 nm in films, which is significantly narrowed in comparison with 1.63 eV for its isomer **FPIC5**. As a result, **FPIC6** possesses complementary absorption spectra to **PTB7-Th** (Fig. 1a), which is beneficial for higher photocurrent.

2.3. Electrochemical properties and theoretical calculations

The electrochemical properties of **FPIC6** were investigated by cyclic voltammetry (CV). As shown in Fig. S3,[†] **FPIC6** possesses the onset oxidation potential ($E_{\text{ox}}^{\text{on}}$) of 0.64 V, corresponding to the HOMO energy level at -5.44 eV , while the onset reduction potential ($E_{\text{red}}^{\text{on}}$) of **FPIC6** is -0.89 V , corresponding to the LUMO energy level of -3.91 eV . Compared with **FPIC5**, the reduced bandgap of **FPIC6** mainly results from the more significantly lifted HOMO energy level rather than the slightly increased LUMO level, which is a very favourable way for highly efficient solar cells without sacrificing open-circuit voltage (V_{OC}). Meanwhile, the energy levels of **FPIC6** exhibited a good match with the well-known polymer donor **PTB7-Th** (Fig. 1c).

Density functional theory (DFT) calculations at the B3LYP/6-31G(d,p) level were performed to study the geometric and electronic properties of **FPIC6** and **FPIC5** (Fig. S4 and S5[†]). Unlike coplanar **FPIC5**, **FPIC6** possesses a twisted molecular conformation, with the dihedral angles of $\sim 10^\circ$ between the core unit and terminal 2FIC units, which could nicely contribute to its good solubility in organic solvents. In comparison with **FPIC5**, the HOMO orbital of **FPIC6** is further

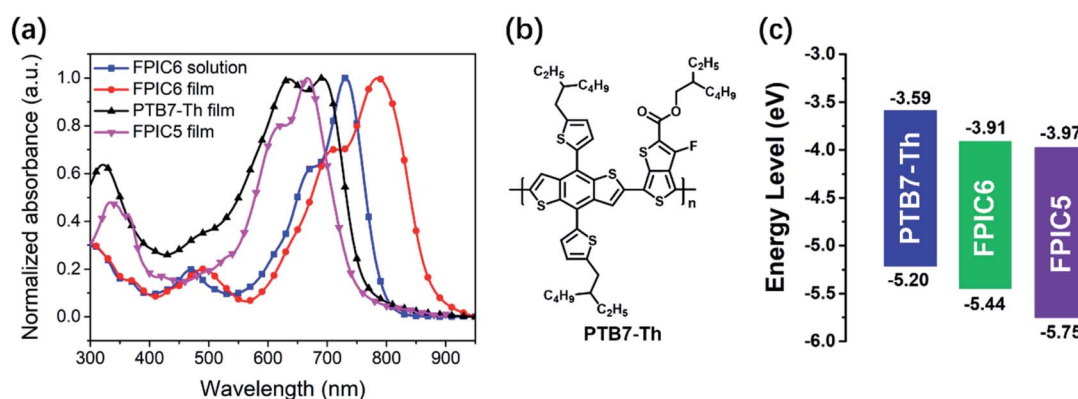


Fig. 1 Absorption spectra of **FPIC6**, **PTB7-Th** and **FPIC5** (a), chemical structure of **PTB7-Th** (b), and energy level diagrams of **FPIC6**, **PTB7-Th** and **FPIC5** (c).



Table 1 Optical and electrochemical properties of FPIC6 and FPIC5

Acceptor	λ_{sol} [nm]	λ_{film} [nm]	λ_{onset} [nm]	$E_{\text{g}}^{\text{optb}}$ [eV]	HOMO ^c [eV]	LUMO ^d [eV]	HOMO ^e [eV]	LUMO ^e [eV]
FPIC6	723	785	871	1.42	-5.44	-3.91	-5.37	-3.44
FPIC5^a	645	666	760	1.63	-5.75	-3.97	-5.61	-3.39

^a These data are from ref. 29. ^b $E_{\text{g}}^{\text{opt}} = 1240/\lambda_{\text{onset}}$. ^c HOMO = $-(E_{\text{ox}}^{\text{on}} - 4.8)$ eV. ^d LUMO = $-(E_{\text{red}}^{\text{on}} - 4.8)$ eV. ^e Calculated from DFT.

localized on the central core units, which satisfactorily follows the design idea of a more fused six-membered ring bridge for this molecule. The LUMO orbital is mainly distributed along the whole molecular backbone, from thieno[3,2-*b*]thiophene-fused pyrene as a core to the terminal electron-withdrawing units, suggesting a more effective intramolecular charge transfer effect in **FPIC6**.⁴⁵ This observation agrees well with the report, in which it is described that the connection through 1,8-positions (indicated by the blue arrow in the structure of **FPIC6** in Scheme 1) of pyrene has much higher conductance than that through 2,9-positions (indicated by the blue arrow in the structure of **FPIC5** in Scheme 1) of it.⁴⁶ Therefore, the improved connectivity in **FPIC6** can explain its extended conjugation, enhanced ICT effect, and narrower bandgap as well as red-shifted absorption spectrum. Moreover, the calculated HOMO/LUMO energy levels are $-5.37/-3.44$ eV for **FPIC6** and $-5.61/-3.39$ eV for **FPIC5**. **FPIC6** showed a similar LUMO level to **FPIC5**, but a 0.24 eV higher HOMO level and a narrower bandgap. The trend is consistent with the CV measurement results.

2.4. Photovoltaic performance

Inverted solar cells with the structure of ITO/ZnO/PFN-Br/**PTB7-Th:FPIC6**/MoO₃/Al were fabricated to investigate the photovoltaic performance of **FPIC6**, and a well-known D-A polymer **PTB7-Th** was chosen as the donor material for the reasons mentioned above. As shown in Table S2,[†] diphenylethane (DPE), 1,8-diiodooctane (DIO), and especially 1-chloronaphthalene (CN) were used as processing additives to optimize the microstructure morphology as well as the photovoltaic performance of **PTB7-Th:FPIC6** cells. Without CN, the solar cell just

achieved a reasonable PCE of 9.11%, with a V_{OC} of 0.79 V, a J_{SC} of 20.16 mA cm⁻², and a fill factor (FF) of 0.58 (Fig. 2a and Table 2). The cells processed with 1% CN (vol/vol) in chloroform have achieved the highest PCE of 11.55%, with a V_{OC} of 0.79 V, a J_{SC} of 21.50 mA cm⁻², and a FF of 0.67. On the other hand, the isomer **FPIC5** based OSCs with the same donor **PTB7-Th** exhibited a moderate PCE of 8.45% with a similar V_{OC} of 0.755 V, but a much lower J_{SC} of 15.3 mA cm⁻².³² According to the formula $E_{\text{loss}} = E_{\text{g}}^{\text{opt}} - eV_{\text{OC}}$,^{47,48} the corresponding energy loss (E_{loss}) of **FPIC6** is 0.63 eV, which is much lower than that of **FPIC5** (0.88 eV). The enhanced photovoltaic performance of the optimized devices can be derived from the simultaneously enhanced J_{SC} and FF. As shown in Fig. 2b, by adding CN as the processing additive, the external quantum efficiency (EQE) responses were improved in the range of 400–760 nm. The current density integrated from the EQE curve of the champion cells processed with 1% CN is 20.53 mA cm⁻², which agrees well with the J_{SC} value from J - V measurements. The much higher and broader EQEs can well explain their much higher photocurrent as compared to **PTB7-Th:FPIC5** cells with a J_{SC} of 15.1 mA cm⁻². Moreover, **PTB7-Th:FPIC6** cells with 0.5–1.5% CN exhibited an ~25 nm blue-shift of EQE when compared to that of the devices without CN. This phenomenon is far less common in OSCs. These blue-shifted EQEs may originate from the reduced packing/aggregation of **FPIC6** in the blend film with 1% CN.^{49–52}

2.5. Molecular orientation and crystallinity

To confirm this hypothesis, the absorption of **PTB7-Th:FPIC6** blend films was also recorded. As shown in Fig. S6,[†] the **PTB7-**

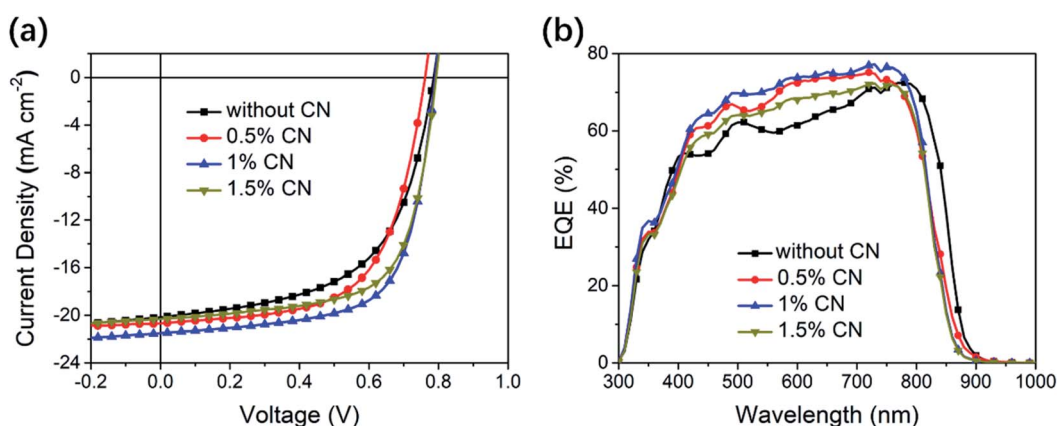


Fig. 2 J - V curves (a) and EQE spectra (b) of OSCs based on **PTB7-Th:FPIC6**.



Table 2 Photovoltaic performance of the solar cells based on PTB7-Th:FPIC6

CN [vol%]	V_{OC} [V]	J_{SC} [mA cm ⁻²]	FF	PCE ^a [%]	J_{EQE} [mA cm ⁻²]
0	0.79 (0.78 ± 0.01)	20.16 (20.14 ± 0.02)	0.58 (0.58 ± 0.02)	9.11 (9.10 ± 0.01)	19.43
0.5	0.76 (0.75 ± 0.01)	20.68 (20.65 ± 0.03)	0.62 (0.60 ± 0.02)	9.91 (9.89 ± 0.02)	20.05
1	0.79 (0.78 ± 0.01)	21.50 (21.46 ± 0.04)	0.67 (0.66 ± 0.01)	11.55 (11.52 ± 0.03)	20.53
1.5	0.79 (0.78 ± 0.01)	20.31 (20.28 ± 0.03)	0.67 (0.64 ± 0.03)	10.75 (10.74 ± 0.01)	19.12

^a The average values in the parentheses are from 20 devices.

Th:FPIC6 blend film cast from its chloroform solution without CN showed an overall blue-shifted absorption compared to that of the neat **FPIC6** film. The shoulder peak at 775 nm was plainly weakened, revealing the interfered π - π stacking of **FPIC6** in the presence of **PTB7-Th**, which in turn dramatically enhanced the ICT of **FPIC6** as indicated by the sharp increase of the peak at 680 nm. As for the blend film cast from the solution with 1% CN as an additive, its overall absorption exhibited a larger blue shift of ~ 30 nm than the film cast without involving CN. These changes may result from the confined crystallization of **FPIC6** during the chloroform evaporation in the presence of CN. Our results indicated that CN as an additive can effectively tune the crystallization behaviour of small molecule acceptor **FPIC6**, resulting in optimized morphology.

XRD (Fig. S7[†]) patterns firstly demonstrate an amorphous **PTB7-Th** donor polymer and crystalline **FPIC6**, respectively. Reading the diffractograms of the films of neat **FPIC6** and **PTB7-Th:FPIC6** blends without CN, the same diffraction peak at $2\theta = 7.95^\circ$, corresponding to the same d -spacing of 11.11 Å for the crystal planes of **FPIC6** and its blends with **PTB7-Th** without the additive of CN, can be found, while the crystallinity of **FPIC6** was clearly restricted by **PTB7-Th**. For the blend film with 1% CN, one broad diffraction peak ranging from 6.8–8.3° is well pronounced. This may also confirm that there are both H- and J-aggregation in the blend film with 1% CN.^{52,53}

Grazing incidence wide-angle X-ray scattering (GIWAXS) was also employed to furthermore investigate the molecular orientation of **FPIC6** and **PTB7-Th** in their thin films (Fig. 3). The out-

of-plane (OOP) diffractogram of the **PTB7-Th** neat film shows predominant face-on orientation with π - π stacking distances of 3.99 Å, which agrees well with both an in-plane distinct diffraction peak at around 2.57 nm⁻¹ for (100) and a less pronounced peak at 8.50 nm⁻¹ for (200) representing its sidechain packing. As for the neat film of **FPIC6**, both diffractions at wide angles of *ca.* $q_z \sim 16.7$ nm⁻¹ (corresponding to its π - π stacking distances of 3.75 Å) and ~ 5.0 nm⁻¹ are found in its OOP profile simultaneously, which suggests clearly the mixing of edge-on and face-on orientation of **FPIC6**. In the meantime, it also shows high crystallinity with the (100) diffraction peak at $q_y \sim 3.3$ nm⁻¹ and with clear high order diffraction peaks at 5.2 nm⁻¹ for (200) and 6.6 nm⁻¹ for (300) in the in-plane (IP) direction, revealing long-range order of the lamellar packing of **FPIC6**. For the blend films, either with or without CN as an additive, dominating face-on orientation was preserved with significantly reduced intensity of (100) diffraction peaks in the OOP direction. It is noteworthy that the (100) diffraction peak of the blend film with 1% CN shifted to small q values compared to that of the blend film without CN, which might also result from the reduced packing/aggregation. These can well explain the blue shift of EQE and absorption spectra by adding CN.

As shown in Fig. S8,[†] differential scanning calorimetry (DSC) studies revealed a cold crystallization peak at 217 °C and a melting peak at 251 °C for **FPIC6** in the heating process, and two crystallization peaks at 231 °C and 235 °C upon cooling, confirming its crystalline structure.^{54–57} In contrast, when it is blended with amorphous **PTB7-Th**, the crystallization of **FPIC6**

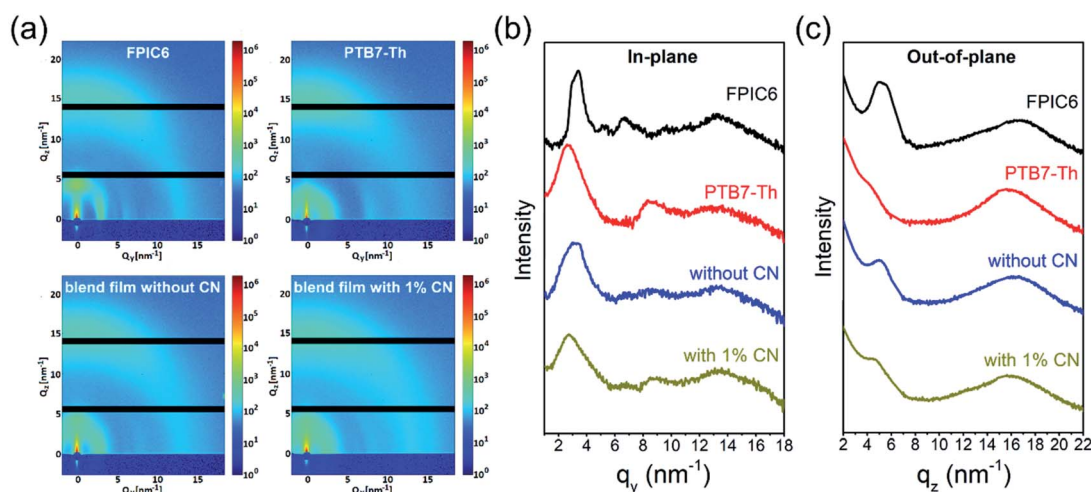


Fig. 3 GIWAXS images (a), in-plane (b), and out-of-plane (c) scattering profiles for neat and PTB7-Th:FPIC6 blend films.



turned to be remarkably restricted either in the absence or presence of CN, due to the strong interference of PTB7-Th on the π - π stacking of FPIC6 as discussed above.

2.6. PL quenching and exciton lifetimes

Correspondingly, the PTB7-Th:FPIC6 blend film cast from chloroform with 1% CN also showed higher photoluminescence (PL) quenching efficiency compared to the blend film evaporated without CN (Fig. S9[†]), regardless of the excitation wavelength located at the absorption maximum of PTB7-Th or FPIC6, indicating that the active layer with 1% CN has more efficient photo-induced exciton dissociation and charge transport behaviour. The exciton dissociation was studied by time-resolved photoluminescence (TR-PL) measurements with photoexcitation at an excitation wavelength of 780 nm, whose photon flux is 8.3×10^{12} photons per cm^2 . The monitored emission is at the wavelength integrated from 820 to 880 nm. As shown in Fig. 4, the blend films show much shorter exciton lifetimes (τ) compared to that of neat FPIC6 films (134 ps), due to the photoinduced charge transfer in the excited state.⁵⁸

2.7. Charge carrier mobilities and morphology

Hole and electron mobilities of PTB7-Th:FPIC6 blend films were measured by the space-charge limited current (SCLC) method (Fig. S10 and Table S3[†]). The PTB7-Th:FPIC6 blend film with 1% CN shows an electron mobility of $6.2 \times 10^{-5} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$, which is twice higher than that of the blend film without CN. Thus, the active layer cast with 1% CN exhibited more balanced charge carrier transport and enhanced FF as mentioned above.

In order to understand the effect of the addition of 1% CN, the surface and bulk morphologies of the blend films were also studied by atomic force microscopy (AFM) and transmission electron microscopy (TEM). AFM images in Fig. 5 and S11[†] indicate that the film without CN shows tiny FPIC6 crystals in the film with a large RMS roughness of 7.09 nm, while the film with 1% CN shows fine and smooth structures, and the RMS roughness was also reduced to 1.03 nm. In the TEM images shown in Fig. 5, unlike the film without 1% CN exhibiting large

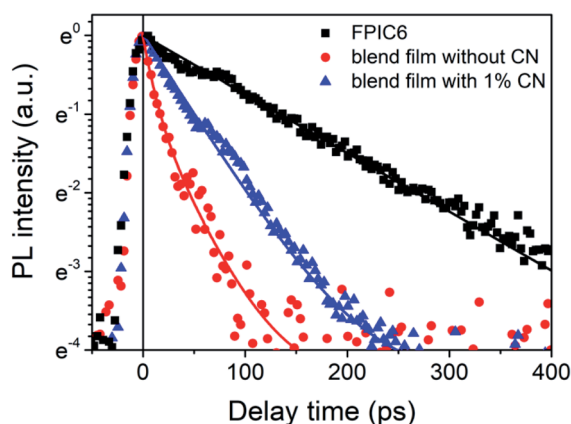


Fig. 4 Time-resolved PL decay spectra of FPIC6 and PTB7-Th:FPIC6 blend films.

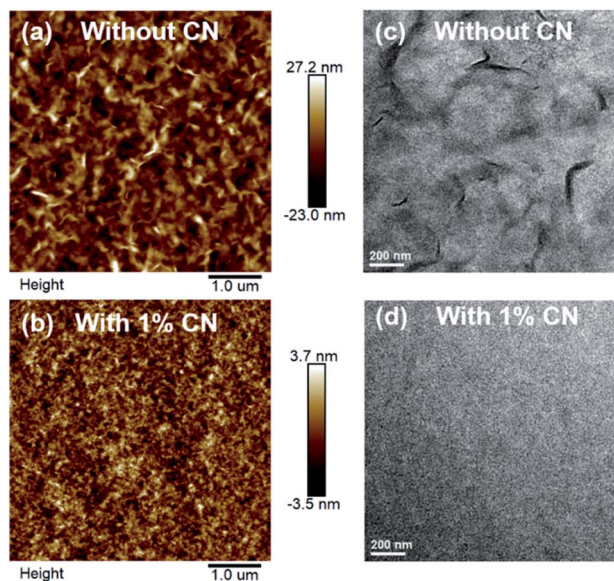


Fig. 5 (a and b) AFM height images and (c and d) TEM images of PTB7-Th:FPIC6 blend films without CN (above) and with 1% CN (bottom).

domain features, the film with 1% CN was quite uniform and exhibited good miscibility, which agrees quite well with the AFM images. These results indicate that the addition of 1% CN can reduce large crystals from FPIC6 and increase the miscibility of the two components, which is beneficial for increasing D/A interface areas and effective exciton dissociation, leading to much higher J_{SC} and FF.

3. Experimental section

All reagents were purchased from Alfa Aesar Co., Aladdin Co., Derthon Co. and other commercial suppliers. 1,6-Bis(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)pyrene (1) and ethyl 2-bromothieno[3,2-*b*]thiophene-3-carboxylate were prepared according to the literature.^{35,59}

3.1. Synthesis of diethyl 2,2'-(pyrene-1,6-diyl)bis(thieno[3,2-*b*]thiophene-3-carboxylate) (2)

A mixture of ethyl 2-bromothieno[3,2-*b*]thiophene-3-carboxylate (1.18 g, 4.05 mmol) and compound 1 (735 mg, 1.62 mmol) in a mixed solution of toluene (35 mL)/ethanol (15 mL)/ K_2CO_3 aqueous solution (2 M, 15 mL) was degassed before $\text{Pd}(\text{PPh}_3)_4$ (100 mg, 0.09 mmol) was added. The mixture was heated to reflux for 24 h, and then cooled to room temperature. After filtration, compound 2 was obtained as a yellow solid (894 mg, 88.7%). ^1H NMR (500 MHz, CDCl_3 , δ/ppm): 8.22 (d, $J = 7.1$ Hz, 2H), 8.10–8.06 (m, 4H), 7.57 (d, $J = 5.0$ Hz, 2H), 7.38 (d, $J = 5.0$ Hz, 2H), 7.19 (d, $J = 7.5$ Hz, 2H), 3.98 (br, 4H), 0.68 (br, 6H).

3.2. Synthesis of compound 4

To a solution of 1-bromo-4-hexylbenzene (2.07 g, 8.58 mmol) in 50 mL of THF at -78 °C was added *n*-BuLi (3.4 mL, 2.5 M in hexane) under nitrogen. The mixture was kept at -78 °C for 1 h,



and then compound **2** (894 mg, 1.44 mmol) was added. The mixture was stirred at room temperature overnight and then poured into water and extracted twice with ethyl acetate. After removing the solvent, the crude compound **3** was dissolved in 50 mL dry toluene and 1.0 g Amberlyst 15 was added, and then the mixture was refluxed for 6 h under nitrogen. After cooling to room temperature, the mixture was filtered and the solvent was removed. The crude product was purified by column chromatography on silica gel using a mixture of petroleum ether/dichloromethane (10 : 1, v/v) as the eluent to give compound **4** as an orange solid (386 mg, 23.5%). ¹H NMR (400 MHz, CDCl₃, δ/ppm): 8.02 (d, *J* = 8.0 Hz, 2H), 7.96 (d, *J* = 8.1 Hz, 2H), 7.88 (s, 2H), 7.22–7.17 (m, 12H), 7.03 (d, *J* = 8.3 Hz, 8H), 2.56–2.49 (m, 8H), 1.56–1.51 (m, 8H), 1.33–1.25 (m, 24H), 0.88–0.83 (m, 12H). ¹³C NMR (125 MHz, CDCl₃, δ/ppm): 142.35, 141.23, 141.04, 137.44, 136.69, 135.05, 130.37, 129.49, 129.32, 127.78, 127.55, 125.87, 125.39, 123.29, 120.21, 118.65, 57.99, 34.96, 31.21, 30.72, 29.23, 28.49, 22.11, 13.60. MS (MALDI-TOF, *m/z*): 1142.605.

3.3. Synthesis of compound **5**

To a solution of compound **4** (150 mg, 0.13 mmol), 2 mL of DMF, and 30 mL of 1,2-dichloroethane at 0 °C was added POCl₃ (0.24 mL, 2.62 mmol) under nitrogen. The mixture was kept at 0 °C for 0.5 h, and then was heated to 85 °C for 24 h. After cooling to room temperature, 100 mL NaHCO₃ aqueous solution was added and the mixture was extracted with dichloromethane. After removal of the solvent, the crude product was purified by column chromatography on silica gel using a mixture of petroleum ether/dichloromethane (2 : 1, v/v) as the eluent to give compound **5** as a red solid (132 mg, 84.1%). ¹H NMR (500 MHz, CDCl₃, δ/ppm): 9.85 (s, 2H), 8.09 (d, *J* = 8.0 Hz, 2H), 8.04 (d, *J* = 8.1 Hz, 2H), 7.88 (s, 2H), 7.85 (s, 2H), 7.16 (d, *J* = 8.3 Hz, 8H), 7.05 (d, *J* = 8.3 Hz, 8H), 2.61–2.50 (m, 8H), 1.59–1.53 (m, 8H), 1.36–1.22 (m, 24H), 0.85 (t, *J* = 6.7 Hz, 12H). MS (MALDI-TOF, *m/z*): 1199.473.

3.4. Synthesis of FPIC6

To a solution of compound **5** (164 mg, 0.14 mmol), 2-(5,6-difluoro-3-oxo-2,3-dihydro-1*H*-inden-1-ylidene)malononitrile (2FIC, 158 mg, 0.68 mmol), and 50 mL chloroform was added 1 mL pyridine under nitrogen. The mixture was heated to reflux for 16 h. After cooling to room temperature, the mixture was poured into 150 mL methanol and filtered. The crude product was purified by column chromatography on silica gel using a mixture of petroleum ether/dichloromethane (1 : 3, v/v) as the eluent to give FPIC6 as a black solid (171 mg, 77.0%). ¹H NMR (400 MHz, CDCl₃, δ/ppm): 8.71 (s, 2H), 8.51–8.47 (m, 2H), 8.17 (s, 2H), 8.13 (d, *J* = 8.1 Hz, 2H), 8.07 (d, *J* = 8.2 Hz, 2H), 7.97 (s, 2H), 7.61 (m, 2H), 7.24–7.20 (m, 8H), 7.12–7.10 (m, 8H), 2.60–2.53 (m, 8H), 1.63–1.52 (m, 8H), 1.31–1.20 (m, 24H), 0.84–0.77 (m, 12H). IR: $\nu = 2216\text{ cm}^{-1}$ (C≡N), 1700 cm^{-1} (C=O). MS (MALDI-TOF, *m/z*): 1622.687.

4. Conclusions

In summary, using pyrene fused with two thieno[3,2-*b*]thiophenes *via* its [*f,g,m-n*] bay-position (1,8-position) as six-

membered ring bridges as the core, and 2-(5,6-difluoro-3-oxo-2,3-dihydro-1*H*-inden-1-ylidene)malononitrile (2FIC) as the end-groups, a non-coplanar small molecule acceptor FPIC6 has been designed and successfully synthesized. Compared with its isomer FPIC5 with five-membered ring bridges *via* the [*a,h*] ortho-position at 2,9-positions, FPIC6 with a more fused core and improved connectivity exhibits a 119 nm red-shifted absorption peak in films and a much narrower E_g^{opt} of 1.42 eV. FPIC6 possesses a slightly higher LUMO energy level (−3.91 eV), but a much higher HOMO energy level of −5.44 eV than those of its structural isomer FPIC5. By adding 1% CN as the processing additive, the PCE of PTB7-Th:FPIC6 solar cells was improved from 9.11% to 11.55%, with a V_{OC} of 0.79 V, a J_{SC} of 21.50 mA cm^{−2}, and a FF of 0.67. It is obvious that the better photoluminescence quenching behaviour, more balanced charge carrier mobilities, and optimal morphology are responsible for the much better photovoltaic performance due to the use of CN as the additive. In contrast, solar cells based on PTB7-Th:FPIC5 just achieved a PCE of 8.45%. Our results clearly confirm such 1,8-position connectivity on the pyrene core as the crucial origin of the lowered bandgap and the improved photocurrent and PCE for FPIC6. This strongly indicates that such position dependent connectivity can be a key-design strategy for future A–D–A type narrow bandgap NFAs in common.

Conflicts of interest

There are no conflicts to declare.

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Notes and references

- 1 Y. Lin, J. Wang, Z.-G. Zhang, H. Bai, Y. Li, D. Zhu and X. Zhan, *Adv. Mater.*, 2015, **27**, 1170.
- 2 Y. Lin, Z.-G. Zhang, H. Bai, J. Wang, Y. Yao, Y. Li, D. Zhu and X. Zhan, *Energy Environ. Sci.*, 2015, **8**, 610.
- 3 W. Zhao, S. Li, H. Yao, S. Zhang, Y. Zhang, B. Yang and J. Hou, *J. Am. Chem. Soc.*, 2017, **139**, 7148.
- 4 Y. Wang, Y. Zhang, N. Qiu, H. Feng, H. Gao, B. Kan, Y. Ma, C. Li, X. Wan and Y. Chen, *Adv. Energy Mater.*, 2018, **8**, 1702870.
- 5 Z. Luo, H. Bin, T. Liu, Z.-G. Zhang, Y. Yang, C. Zhong, B. Qiu, G. Li, W. Gao, D. Xie, K. Wu, Y. Sun, F. Liu, Y. Li and C. Yang, *Adv. Mater.*, 2018, **30**, 1706124.
- 6 J. Wang, J. Zhang, Y. Xiao, T. Xiao, R. Zhu, C. Yan, Y. Fu, G. Lu, X. Lu, S. R. Marder and X. Zhan, *J. Am. Chem. Soc.*, 2018, **140**, 9140.
- 7 Y. Yang, Z.-G. Zhang, H. Bin, S. Chen, L. Gao, L. Xue, C. Yang and Y. Li, *J. Am. Chem. Soc.*, 2016, **138**, 15011.
- 8 C. Yan, S. Barlow, Z. Wang, H. Yan, A. K.-Y. Jen, S. R. Marder and X. Zhan, *Nat. Rev. Mater.*, 2018, **3**, 18003.
- 9 P. Cheng, G. Li, X. Zhan and Y. Yang, *Nat. Photonics*, 2018, **12**, 131.
- 10 T. Li, S. Dai, Z. Ke, L. Yang, J. Wang, C. Yan, W. Ma and X. Zhan, *Adv. Mater.*, 2018, **30**, 1705969.
- 11 R. Yu, H. Yao and J. Hou, *Adv. Energy Mater.*, 2018, **8**, 1702814.
- 12 D. He, F. Zhao, L. Jiang and C. Wang, *J. Mater. Chem. A*, 2018, **6**, 8839.
- 13 Z. Xiao, S. Yang, Z. Yang, J. Yang, H.-L. Yip, F. Zhang, F. He, T. Wang, J. Wang, Y. Yuan, H. Yang, M. Wang and L. Ding, *Adv. Mater.*, 2019, **31**, 1804790.
- 14 J. Yuan, Y. Zhang, L. Zhou, G. Zhang, H.-L. Yip, T.-K. Lau, X. Lu, C. Zhu, H. Peng, P. A. Johnson, M. Leclerc, Y. Cao, J. Ulanski, Y. Li and Y. Zou, *Joule*, 2019, **3**, 1140.
- 15 H. Wu, H. Fan, S. Xu, L. Ye, Y. Guo, Y. Yi, H. Ade and X. Zhu, *Small*, 2019, **15**, 1804271.
- 16 Y. Li, L. Zhong, J.-D. Lin, F.-P. Wu, H.-J. Bin, Z. Zhang, L. Xu, Z.-Q. Jiang, Z.-G. Zhang, F. Liu, T. P. Russell, Y. Li, L.-S. Liao and S. R. Forrest, *Sol. RRL*, 2017, **1**, 1700107.
- 17 F. Liu, Z. Zhou, C. Zhang, T. Vergote, H. Fan, F. Liu and X. Zhu, *J. Am. Chem. Soc.*, 2016, **138**, 15523.
- 18 B. Fan, D. Zhang, M. Li, W. Zhong, Z. Zeng, L. Ying, F. Huang and Y. Cao, *Sci. China: Chem.*, 2019, **62**, 746.
- 19 Q. Liu, Y. Jiang, K. Jin, J. Qin, J. Xu, W. Li, J. Xiong, J. Liu, Z. Xiao, K. Sun, S. Yang, X. Zhang and L. Ding, *Sci. Bull.*, 2020, **65**, 272.
- 20 Y. Chang, T.-K. Lau, M.-A. Pan, X. Lu, H. Yan and C. Zhan, *Mater. Horiz.*, 2019, **6**, 2094.
- 21 L. Meng, Y. Zhang, X. Wan, C. Li, X. Zhang, Y. Wang, X. Ke, Z. Xiao, L. Ding, R. Xia, H.-L. Yip, Y. Cao and Y. Chen, *Science*, 2018, **361**, 1094.
- 22 O. P. Lee, A. T. Yiu, P. M. Beaujuge, C. H. Woo, T. W. Holcombe, J. E. Millstone, J. D. Douglas, M. S. Chen and J. M. J. Fréchet, *Adv. Mater.*, 2011, **23**, 5359.
- 23 H. Huang, L. Xiao, B. Yang, T. Lai, L. Zhang, K. Gao, X. Peng and Y. Cao, *RSC Adv.*, 2016, **6**, 59218.
- 24 S.-S. Li, K.-J. Jiang, C.-C. Yu, J.-H. Huang, L.-M. Yang and Y.-L. Song, *New J. Chem.*, 2014, **38**, 4404.
- 25 D. Yu, Y. Liu, M. Xiao, Q. Fan, W. Su, X. Li, H. Tan, Y. Wang, R. Yang and W. Zhu, *Dyes Pigm.*, 2016, **125**, 151.
- 26 K. Oniwa, H. Kikuchi, H. Shimotani, S. Ikeda, N. Asao, Y. Yamamoto, K. Tanigaki and T. Jin, *Chem. Commun.*, 2016, **52**, 4800.
- 27 Y. Gong, X. Zhan, Q. Li and Z. Li, *Sci. China: Chem.*, 2016, **59**, 1623.
- 28 T. M. Figueira-Duarte and K. Müllen, *Chem. Rev.*, 2011, **111**, 7260.
- 29 Q. Fan, Y. Liu, M. Xiao, H. Tan, Y. Wang, W. Su, D. Yu, R. Yang and W. Zhu, *Org. Electron.*, 2014, **15**, 3375.
- 30 N. J. Jeon, J. Lee, J. H. Noh, M. K. Nazeeruddin, M. Gratzel and S. I. Seok, *J. Am. Chem. Soc.*, 2013, **135**, 19087.
- 31 J.-H. Kim, H. U. Kim, I.-N. Kang, S. K. Lee, S.-J. Moon, W. S. Shin and D.-H. Hwang, *Macromolecules*, 2012, **45**, 8628.
- 32 G. Cai, P. Xue, Z. Chen, T. Li, K. Liu, W. Ma, J. Lian, P. Zeng, Y. Wang, R. P. S. Han and X. Zhan, *Chem. Mater.*, 2019, **31**, 6484.
- 33 J. Wu, Y. Chen, B. Hu, Z. Pang, Z. Lu and Y. Huang, *Chem.–Eur. J.*, 2019, **25**, 6385.
- 34 Y. Wang, B. Liu, C. W. Koh, X. Zhou, H. Sun, J. Yu, K. Yang, H. Wang, Q. Liao, H. Y. Woo and X. Guo, *Adv. Energy Mater.*, 2019, **9**, 1803976.
- 35 X. Wang, L. Feng and L. Zhang, *Dyes Pigm.*, 2013, **97**, 318.
- 36 X. Xu, Z. Li, Z. Bi, T. Yu, W. Ma, K. Feng, Y. Li and Q. Peng, *Adv. Mater.*, 2018, **30**, 1800737.
- 37 Y. Liu, Z. Zhang, S. Feng, M. Li, L. Wu, R. Hou, X. Xu, X. Chen and Z. Bo, *J. Am. Chem. Soc.*, 2017, **139**, 3356.
- 38 Z. Xiao, X. Jia and L. Ding, *Sci. Bull.*, 2017, **62**, 1562.
- 39 Y. Li, N. Zheng, L. Yu, S. Wen, C. Gao, M. Sun and R. Yang, *Adv. Mater.*, 2019, **31**, 1807832.
- 40 J. Yu, J. Cao, H. Tan, W. Peng, Y. Wang and W. Zhu, *Dyes Pigm.*, 2017, **141**, 21.
- 41 W. Peng, G. Zhang, L. Shao, C. Ma, B. Zhang, W. Chi, Q. Peng and W. Zhu, *J. Mater. Chem. A*, 2018, **6**, 24267.
- 42 J. Sun, X. Ma, Z. Zhang, J. Yu, J. Zhou, X. Yin, L. Yang, R. Geng, R. Zhu, F. Zhang and W. Tang, *Adv. Mater.*, 2018, **30**, 1707150.
- 43 W. Hu, Z. Zhang, Y. Xu, J. Cao, S. Liu, Z. Tang, B. Zhao and J. Zhang, *J. Mater. Chem. C*, 2018, **6**, 6911.
- 44 J. Cao, C. Zuo, J. Yu and Z. Tang, *Dyes Pigm.*, 2017, **139**, 201.
- 45 X. Shi, L. Zuo, S. B. Jo, K. Gao, F. Lin, F. Liu and A. K.-Y. Jen, *Chem. Mater.*, 2017, **29**, 8369.
- 46 C. J. Lambert and S.-X. Liu, *Chem.–Eur. J.*, 2018, **24**, 4193.
- 47 D. Veldman, S. C. J. Meskers and R. A. J. Janssen, *Adv. Funct. Mater.*, 2009, **19**, 1939.
- 48 Z. Yao, X. Liao, K. Gao, F. Lin, X. Xu, X. Shi, L. Zuo, F. Liu, Y. Chen and A. K.-Y. Jen, *J. Am. Chem. Soc.*, 2018, **140**, 2054.
- 49 W.-Y. Tan, K. Gao, J. Zhang, L.-L. Chen, S.-P. Wu, X.-F. Jiang, X.-B. Peng, Q. Hu, F. Liu, H.-B. Wu, Y. Cao and X.-H. Zhu, *Sol. RRL*, 2017, **1**, 1600003.



- 50 M. Más-Montoya and R. A. J. Janssen, *Adv. Funct. Mater.*, 2017, **27**, 1605779.
- 51 S. ong Kim, T. K. An, J. Chen, I. Kang, S. H. Kang, D. S. Chung, C. E. Park, Y.-H. Kim and S.-K. Kwon, *Adv. Funct. Mater.*, 2011, **21**, 1616.
- 52 Q. Zhao, J. Liu, H. Wang, M. Li, K. Zhou, H. Yang and Y. Han, *J. Mater. Chem. C*, 2015, **3**, 8183.
- 53 Q. Zhao, X. Yu, J. Liu, Z. Xie and Y. Han, *Org. Electron.*, 2016, **37**, 6.
- 54 S. Holliday, R. S. Ashraf, A. Wadsworth, D. Baran, S. A. Yousaf, C. B. Nielsen, C.-H. Tan, S. D. Dimitrov, Z. Shang, N. Gasparini, M. Alamoudi, F. Laquai, C. J. Brabec, A. Salleo, J. R. Durrant and I. McCulloch, *Nat. Commun.*, 2016, **7**, 11585.
- 55 Y. Li, N. Zheng, L. Yu, S. Wen, C. Gao, M. Sun and R. Yang, *Adv. Mater.*, 2019, **31**, 1807832.
- 56 T. J. Aldrich, M. Matta, W. Zhu, S. M. Swick, C. L. Stern, G. C. Schatz, A. Facchetti, F. S. Melkonyan and T. J. Marks, *J. Am. Chem. Soc.*, 2019, **141**, 3274.
- 57 P. Yin, T. Zheng, Y. Wu, G. Liu, Z.-G. Zhang, C. Cui, Y. Li and P. Shen, *J. Mater. Chem. A*, 2018, **6**, 20313.
- 58 Z. Li, W. Zhang, X. Xu, Z. Genene, D. Di Carlo Rasi, W. Mammo, A. Yartsev, M. R. Andersson, R. A. J. Janssen and E. Wang, *Adv. Energy Mater.*, 2017, **7**, 1602722.
- 59 D. He, L. Qian and L. Ding, *Polym. Chem.*, 2016, **7**, 2329.

